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(54) **METHOD FOR PREPARING A SUBSTRATE WITH MICROMETALLIC MASSES ALIGNED ON THE SURFACE**

VERFAHREN ZUR HERSTELLUNG EINES SUBSTRATS MIT AUF DER OBERFLÄCHE ANGEORDNETEN MIKROMETALLISCHEN MASSES

PROCÉDÉ DE PRÉPARATION D'UN SUBSTRAT AVEC DES MASSES MICROMÉTALLIQUES ALIGNÉES SUR LA SURFACE

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DescriptionField of the Invention

5 **[0001]** The present invention relates to a method for preparing a substrate with micrometallic masses on its surface wherein the size and shape of the micrometallic masses are controlled.

Prior Art

10 **[0002]** Metal nanoparticles with particle diameters of from several nanometers to several tens of nanometers have unusual physical and chemical properties, which are not observed in bulk size due to quantum size effects. When such metal nanoparticles are aligned in an orderly manner, aggregates with new functions hitherto not associated with individual particles are realized. Such aggregates have potential utility in plasmonics elements such as nano optical antennas as well as optical nano circuits. That is, the technology used to prepare and organize metal nanostructures, the size and
15 shape of which are controlled, is the foundation for the nano electronics and plasmonics.

[0003] As the technology used to prepare such fine structures, top-down processes to prepare nanostructures using semiconductor processing technologies such as photo lithography, electron beam exposure, X-ray exposure and the like and bottom-up processes using a metal nanoparticle self organization process (Patent reference 1 and Non-patent reference 1), template methods in which orderly structures voluntarily formed by organic and inorganic materials are cast and the like are available. As far as examples of the latter are concerned, methods in which aluminum oxide cathodes
20 (Patent reference 2, Non-patent references 2 and 3), thin amorphous carbon films (Non-patent reference 4), sodium chloride single crystals (Patent reference 3 and Non-patent reference 5), biopolymers (Non-patent references 6 and 7) and the like are used as casting forms have been reported.

[0004]

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Patent reference 1: Japanese Patent Application Public Disclosure No. 2005-225887

Patent reference 2: Japanese Patent Application Public Disclosure No. 2005-195440

Patent reference 3: Japanese Patent Application Public Disclosure No. 2004-098246

Non-patent reference 1: Adv. Mater. 2001, 13, 1699-1701

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Non-patent reference 2: Chem. Eur. J. 1997, 3, 1951-1956

Non-patent reference 3: Chem. Lett. 2005, 34, 508-509

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Non-patent reference 5: J. Am. Chem. Soc. 2002, 124, 4210-4211

Non-patent reference 6: J.E. Adv. Mater. 2001, 13, 109-113

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Non-patent reference 7: Angew. Chem. Int. Ed. 2002, 38, 1034-1054

[0005] Non-patent document XP 002610248 ("Block Copolymer-Templated Chemistry on Si, Ge, InP, and GaAs surfaces") discloses a process for patterning surface using a block copolymer. According to the second deposition method described in said document, a block copolymer solution is spin-coated onto a wafer to produce a monolayer of
40 block copolymer micelles. This substrate is then immersed in a metal ion solution in the presence of HF that is required to facilitate reduction of the metal ion. Then, the polymer is removed with brief ultrasonication.

Problems to be solved by the Invention

45 **[0006]** Metal nanostructures in which metal masses whose size is in the deca nano region are aligned on a substrate are extremely difficult to produce inexpensively using a top-down microprocessing technology such as semiconductor processing technologies and the like. In contrast, bottom-up technologies used to integrate molecules and particles have an advantage as far as production cost is concerned since they can be used to prepare large volumes of such structures inexpensively. However, inorganic templates are not easy to prepare and decompose after they are used and organic
50 templates that are more readily available lack dependability due to structural disarray and defects generation. Therefore, the development of an industrial manufacturing technology utilizing an organic template of excellent dependability is urgently needed.

[0007] The objective of the present invention is to present a metal nanostructure having metal masses with zero valence aligned on a substrate surface wherein the size and shape of the metal masses are controlled and a manufacturing
55 process thereof.

Means to Solve the Problems

[0008] A micro phase separation membrane (Japanese Patent Application Public Disclosure No. 2004-124088) formed using an amphiphilic block copolymer developed by a group of the inventors can assume a state containing numerous hydrophilic micro diameter cylinders inside the membrane in the direction perpendicular to the membrane surface. The inventors discovered that metal ions are localized in the hydrophilic micro diameter cylinders in a micro phase separation membrane when the film is formed using a solution containing an amphiphilic block copolymer and desired metal ions or when a micro phase separation membrane of the amphiphilic block copolymer is formed and subsequently brought in contact with a solution containing the metal ion. The metal ion localized in the hydrophilic micro diameter cylinders is reduced and the micro phase separation membrane is removed to form dot-shaped (or sphere-shaped) or pillar-shaped (or cylinder-shaped) metal masses of zero valence on the substrate when a treatment to reduce the micro phase separation membrane and a treatment to remove the organic material including the micro phase separation membrane is conducted (See Examples described later.). The inventors discovered that a substrate containing numerous zero valence metal masses that are aligned at set intervals on the substrate surface was obtained when the metal masses are in the approximate shapes of pillars, dots or combinations thereof with the pillars oriented in a direction approximately perpendicular to the surface. The present invention was completed based on the discovery.

[0009] That is, the invention is a process for preparing a substrate with micrometallic masses aligned on its surface, wherein the substrate contains numerous zero valence fine metal masses aligned on the surface at intervals 10 nm to 50 nm wherein the masses are pillar-shaped or dot-shaped or a combination thereof and have diameters of 3 nm to 15 nm wherein the pillar-shaped masses are oriented in a direction substantially perpendicular to the surface, wherein the process comprises the steps of preparing a micro phase separation membrane of an amphiphilic block copolymer on a substrate wherein the membrane has a thickness of 50nm to 1 μ m and then exposing the membrane to a solution containing metal ions, (b) drying the membrane to form a micro phase separation membrane with the metal ions localized in its hydrophilic segment, and (c) irradiating vacuum ultraviolet light on the micro phase separation membrane under vacuum to reduce the metal ions and, as a result of the irradiation, organic materials including those in the micro phase separation membrane are completely removed, leaving the fine metal masses on the surface.

Advantages of the Invention

[0010] The films for preparing the metal nanostructures of the present invention make it possible to form metal nano dots or nano pillars on a broad range of optional substrates without regard to shape or to the material of construction and present a foundation material and a foundation technology on which to build nano electronics and plasmonics.

Brief Description of the Drawings**[0011]**

Figure 1 shows the method for preparing a substrate of the present invention wherein fine metal masses are lined in an orderly manner on the surface.

Figure 2 shows the mode in which a micro phase separation membrane is formed.

Figure 3 shows an AFM shape image of a micro phase separation membrane in which gold ions are bonded.

Figure 4 shows the FT-IR spectra before and after irradiation of vacuum ultraviolet light.

Figure 5 shows the AFM shape image of the gold nano dots formed on a silicon substrate.

Figure 6 shows the AFM shape image of the Ag nano dots prepared. Images resulting from (a) using a molar ratio of 20 for EO/Ag⁺ and (b) using a molar ratio of 120 for EO/Ag⁺ are shown.

Figure 7 shows a) UV spectrum and b) XPS spectrum of Ag nano dots formed on a quartz glass substrate. a) (A) A membrane treated with Ag⁺ and (B) a sample after vacuum ultraviolet light irradiation are shown. b) (A) A membrane treated with Ag⁺, (B) a sample after vacuum ultraviolet light irradiation and (C) a sample oxidized at 600°C are shown.

Figure 8 shows an AFM shape image of Ag nano dots. a) shows the membrane before a treatment, b) shows the membrane after a treatment, c) shows the topological image after a vacuum ultraviolet light irradiation and d) shows the height image.

Figure 9 shows the AFM height image of the nano phase separation membrane obtained by subjecting an amphiphilic block copolymer, PEO272-b-PMA(Az)74, to an interaction with palladium ion. (a) results from EO/Pd = 16.36, (b) results from EO/Pd = 5.54 and (c) results from EO/Pd = 3.27.

Figure 10 shows an AFM height image of palladium nano dots formed on a silicon substrate.

Figure 11 shows an AFM shape image of the micro phase separation membrane prepared in Example 5.

Figure 12 shows a transmission electron microscope photograph of the micro phase separation membrane prepared in Example 5.

Figure 13 shows an AFM shape image of the micro phase separation membrane prepared in Example 6.

Figure 14 shows the FT-IR spectrum of the nano phase separation membrane prepared in Example 6 and the change in the spectrum before and after vacuum ultraviolet light irradiation.

Figure 15 is a spectrum of the gold nano dots in Example 6 obtained using X-ray electron spectroscopy.

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Detailed Description of the Invention

[0012] The micro phase separation membrane used in the present invention contains numerous approximately circular hydrophilic cylinders with a diameter from 3 nm to 15 nm in the membrane oriented in the direction perpendicular to the membrane surface and spaced at 10 nm to 50 nm intervals and comprises an amphiphilic block copolymer containing localized metal ions in the hydrophilic segments.

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[0013] The amphiphilic block copolymer is a block copolymer obtained by bonding incompatible polymers, a hydrophilic polymer component (A) and a hydrophobic polymer component (B), using a covalent bond, and molecular weight distributions (Mw/Mn) for the hydrophilic polymer component (A) and the hydrophobic polymer component (B) of 1.3 or less are preferred.

15

[0014] Based on the quantitative relationship between the hydrophilic and hydrophobic components, the hydrophilic component is thought to constitute a domain region (cylinder like domain) and the hydrophobic component is thought to constitute a matrix region. The thickness of the membrane formed is ordinarily about 50 nm to 1 μ m.

[0015] The molecular weight distribution (Mw/Mn) refers to a value calculated from the weight average molecular weight, Mw, and the number average molecular weight, Mn, measured using gel permeation chromatography and calculated in reference to polystyrene .

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[0016] The molecular weight of the block copolymer described above is preferably from 5,000 to 100,000 and more preferably from 10,000 to 50,000.

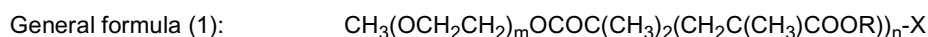
[0017] As the hydrophilic polymer component (A) described above, poly(ethylene oxide), poly(propylene oxide), poly(vinyl alcohol), poly(acrylic acid), poly(methacrylic acid), poly(acrylamide), oligo(ethylene oxide), crown ether and poly(methacrylate) or poly(acrylate) and the like containing cryptand and sugar chains on the side chain, for example, may be cited.

25

[0018] In addition, as the hydrophobic polymer component (B) described above, poly(methacrylate), poly(acrylate), polystyrene, vinyl polymers and the like containing mesogene side chains, long chain alkyl side chains or hydrophobic side chains, for example, may be cited. Mesogene side chains refer to strongly oriented molecular chains such as liquid crystalline molecules.

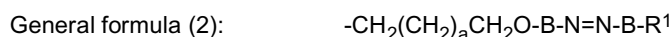
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[0019] Amphiphilic block copolymers represented by the general formula (1) below are preferred.



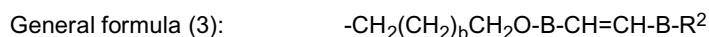
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(In the formula, m and n may be identical to each other or different and are each integers from five to five hundred, R is a substituent represented by the general formulae (2) and (3) shown below and X represents a hydrogen atom or a halogen atom.)



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(In the formula, a is an integer zero to twenty, R¹ is a hydrogen or an alkyl group with one to twenty-two carbon atoms, and B represents a p-phenylene group.)



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(In the formula, b is an integer zero to twenty, R² is a hydrogen or an alkyl group containing one to twenty-two carbon atoms, and B represents a phenylene group.)

[0020] The metal ions used in the present invention are not particularly restricted, but the ions of copper group metals and platinum group metals can be used. As the copper group metals, Ag and Au may be cited, and Pt, Pd, Rh, Ir, Ru and Os may be cited as platinum group metals. They may be used individually or as a mixture thereof. These metal ions are suited for the objective of the invention of this application since they have very low ionization tendency and are chemically stable allowing them to be readily reduced.

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[0021] The metal ions may be formed by dissolving metal precursors in a solvent, and metal salts are preferred as the precursors.

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[0022] Examples of such metal salts are cited below.

[0023] As a gold salt, those containing Au⁺, Au⁺³ or Au⁴⁺ may be cited. Those represented by Au^IX, Au^{III}X₃, M[Au^{IV}X₆] (in the formula, X represents a negative ion, Cl⁻ or CN⁻, and M represents a monovalent cation Na⁺, K⁺ or H⁺) may be

used, and AuCN, AuCl₃, HAuCl₄, NaAuCl₆, KAuCl₆, Au₂O₃ and the like, for example, may be cited.

[0024] As a silver salt, those containing Ag⁺ may be cited. Those represented by Au¹X (in the formula X represents a negative ion, F⁻, Cl⁻, Br⁻, I⁻, CN⁻, SCN⁻, NO₃⁻, PF₆⁻, CO₃⁻, ClO₄⁻, SO₄²⁻, acac⁻, CH₃CO₂⁻, CF₃CO₂⁻, CF₃SO₃⁻, O⁻, O₂⁻, S⁻, PO₄³⁻ or IO₃⁻) may be used, and AgF, AgCl, AgBr, AgI, AgCN, AgSCN, AgNO₃, AgPF₆, Ag₂CO₃, AgClO₄, Ag₂SO₄, acacAg, CH₃CO₂Ag, CF₃CO₂Ag, CF₃SO₃Ag, AgO, Ag₂O, Ag₂S, Ag₃PO₄, AgIO₃, AgSbF₆, AgAsF₆, Ag₈W₄O₁₆ and the like, for example, may be cited.

[0025] As a platinum salt, those containing Pt²⁺, Pt³⁺ and Pt⁴⁺ may be cited. Those represented by PtX₂, PtX₃, PtX₄, [PtA₆]X₂, M¹₂[PtX], M¹₂[PtX₂Y₂], M¹[PtX₃Y], M¹[X₂Y₂], M¹₂[PtX₆] (in the formula x and Y are each negative ions such as F⁻, Cl⁻, Br⁻, I⁻, OH⁻, CN⁻, NO₃⁻, N³⁻, CH₃COO⁻, SCN⁻, acetylacetonate, 1/2 SO₄²⁻, 1/2 CO₃²⁻ and the like, M¹ is a monovalent cation such as K, Na, H and the like, A is NH₃ or an amine) may be used, and PtCl₂, PtBr₂, PtI₂, Pt(CN)₂, Pt(SCN)₂, PtCl₃, PtBr₃, PtI₃, PtF₄, PtCl₄, PtBr₄, PtI₄, K₂[PtCl₂(acac)₂], H₂PtCl₆ and the like, for example, may be cited.

[0026] As a ruthenium salt, those containing Ru²⁺, Ru³⁺ or Ru⁴⁺ may be cited. Those represented by RuX₂, RuX₃, RuX₄, [RuX₆]M¹₃, M¹[RuX₄] (in the formula x is a halogen such as Cl, Br and the like or an anion such as NO₃⁻, SO₄²⁻ and the like and M¹ is a monovalent cation such as K, Na, Rb, Cs, H and the like) may be used, and RuCl₃, (NH₄)₂RuCl₆, Ru(SO₄)₂, RuS₂, RuO₂, RuO₄, Na₂RuO₄, K₂RuO₄ and the like, for example, may be cited.

[0027] As an iridium salt, those containing Ir⁺, Ir²⁺, Ir³⁺ or Ir⁴⁺ may be cited. Those represented by IrX, IrX₂, IrX₃, IrX₄, [IrX₆]M¹₃, M¹[IrX₄] (in the formula X is a halogen such as Cl, Br and the like or an anion such as SO₄²⁻ and M¹ is a monovalent cation such as K, Na, Rb, Cs or H and the like) may be used, and KIr(SO₄)₂, RbIr(SO₄)₂, CsIr(SO₄)₂ and the like, for example, may be cited.

[0028] As a palladium salt, those containing Pd²⁺ may be cited. Those ordinarily represented by Pd-Z₂ (in the formula Z is a halogen such as Cl, Br, I and the like or a salt such as acetate, trifluoroacetate, acetylacetonate, carbonate, perchlorate, nitrate, sulfate, oxide and the like) may be used, and PdCl₂, PdBr₂, PdI₂, Pd(OCOCH₃)₂, Pd(OCOCF₃)₂, PdSO₄, Pd(NO₃)₂, PdO and the like, for example, may be cited.

[0029] As an osmium salt, those containing Os⁺, Os²⁺, Os³⁺ or Os⁴⁺ may be cited. Those represented by OsX, OsX₂, OsX₃, OsX₄, [OsX₆]M¹₃, M¹[OsX₄] (in the formula X is a halogen such as Cl, Br and the like or an anion such as SO₄²⁻ and M¹ is a monovalent cation such as K, Na, Rb, Cs or H and the like) may be used, and OsBr₄, OsO₄, OsCl₄, Kos(SO₄)₂, RbOs(SO₄)₂, CsOs(SO₄)₂ and the like, for example, may be cited.

[0030] As a rhodium salt, those containing Rh³⁺ may be cited. Those represented by RhX₃, Rh₂X₆, [RhA₆]X₃, M¹₃[RhX₆], M¹[RhX₄] (in the formula x is a halogen such as F, Cl and the like or an anion such as CN, SO₄ and the like and M¹ is a monovalent cation such as K, Na, H and the like and A is NH₃ or an amine) may be used, and Rh₂O₃, RhO₂, Rh₂(SO₄)₃, Rh(OH)₃, Rh(NO₃)₃, RhCl₃, RhF₃, RhCN₃, KRh(SO₄)₂, Na₂RhCl₄, NaRh(SO₄)₂, HRh(SO₄)₂ and the like, for example, can be cited.

[0031] Next, a method to manufacture a substrate containing lines of fine metal masses on the surface is described. The substrate can be prepared using two methods.

[0032] A micro phase separation membrane is ordinarily prepared by dissolving the amphiphilic block copolymer described above in a solvent and forming a membrane. In this method not being part of the present invention, however, a precursor of the metal ions described above is dissolved in a solvent, and a solution obtained by dissolving an amphiphilic block copolymer and the metal ions in a solvent are used to form a membrane.

[0033] As a separate method of the present invention, a membrane is formed on a substrate using a solution of an amphiphilic block copolymer to obtain a micro phase separation membrane. The membrane, for example, is annealed in vacuum at 140°C for twenty-four hours to prepare a highly orderly nano phase separation membrane. By impregnating the membrane in a metal salt solution, a micro phase separation membrane containing localized metal ions in the hydrophilic cylinder segments is prepared.

[0034] As the solvent that can dissolve or disperse the metal salt, water; ketones such as acetone and methyl ethyl ketone; esters such as ethyl acetate; alcohols such as methanol and ethanol; and non-protonic polar solvents such as dimethyl formamide, dimethyl sulfoxide, sulfolane, diglyme and hexamethyl phosphoric acid triamide as well as nitromethane, acetonitrile and the like may be cited. The use of water or a mixture of water and a hydrophilic organic solvent such as alcohols and ketones that blend with water is preferred.

[0035] Poly(ethylene glycol) (PEG), polyvinylpyrrolidone (PVP) and the like may also be added as needed to the solution.

[0036] The metal salt concentration ordinarily used is from 0.001% to saturation. When the solvent is water, from 0.01% to 20% is preferred and from 0.1% to 5% is more preferred.

[0037] The amount of metal ions used per block copolymer is set so that the molar ratio (EO/M) for the metal ion (M) to the ethylene oxide monomer (EO) constituting the poly(ethylene oxide) chain of the block copolymer is from one to one thousand and preferably from one hundred to five hundred. Dots or pillars are ideally formed when the ratio is within this range.

[0038] Next, the solution is used to prepare a membrane. When an amphiphilic block copolymer of the present invention is used, the hydrophilic domain and the hydrophobic domain separate through self aggregation. The positioning through self aggregation is preferably promoted by allowing the domains to self aggregate in the solution or by applying heat.

[0039] More specifically, a membrane can be prepared by dropping a solution containing an amphiphilic block copolymer dropwise onto a substrate or by casting the solution into a die. As specific membrane forming processes, spin coating, casting, bar coating, micro gravure and the like may be used.

[0040] Smooth substrates are preferred as the substrate, and mica sheets, silicone wafers and the like, for example, may be used.

[0041] After applying a coating, a membrane is formed under the following conditions. The solvent is removed at room temperature under ambient pressure or reduced pressure. Furthermore, a better ordered nano phase separation membrane can be prepared upon annealing the membrane at ambient or reduced pressure at a temperature of the isotropic point plus no more than 10°C for at least several hours (ordinarily for one to two days).

[0042] When an amphiphilic block copolymer of general formula (1) is used, a membrane or a film about 50 nm to 1 μm thick and about several square meters in size can be obtained. The process is described in Figure 2.

[0043] As a result, the metal ions are localized in the cylinder section (hydrophilic zone) of the micro phase separation membrane. The result is attributed to non-covalent bonding derived from Van der Waals forces, dipole-dipole interactions, static attraction and the like. The bonding force is weak, and the metal ions are not always selectively bonded to the hydrophilic cylinder zones when, for example, the dispersion concentration is too high. Therefore, the metal ion concentration is preferably suitably adjusted so that a desired occupancy ratio (the proportion with which micells are bonded to the domain region) can be attained. Similar considerations are desired for other conditions (impregnation time, temperature, rotations number and the like).

[0044] Next, a micro phase separation membrane obtained in the manner described above may be laminated on a substrate. A substrate that is oxidized with difficulty by ozone (O₃) is preferred, and silicone wafers, glass, metal and the like, for example, may be used as substrates. This operation may be eliminated if a suitable substrate is used in the process described above.

[0045] Next the metal ions localized in the cylinders in a micro phase separation membrane are reduced, and the organic materials including the micro phase separation membrane are removed. The reduction and organic material removal treatments are conducted in one step.

[0046] As the treatment, vacuum ultraviolet light irradiation, is used. Examples of such an ultraviolet light irradiation treatment and of other treatments not being part of the present invention are described below.

(1) Ultraviolet light Irradiation

[0047] The metal ions in a membrane are excited when a sample (a micro phase separation membrane containing metal ions) under ambient pressure is irradiated with ultraviolet light (wavelength: from about 1 nm to 400 nm) and are reduced by the surrounding membrane components.

[0048] The use of vacuum ultraviolet light (wavelength: about 200 nm or shorter) for the ultraviolet light irradiation is preferred since it allows the reduction treatment and the organic material removal treatment to be conducted in one step. The oxidation power of the metal cations excited by the vacuum ultraviolet light oxidation increases, and the ions are thought to be reduced upon receiving electrons from the membrane components in the surrounding areas or upon receiving electrons from decomposition products when the membrane is decomposed upon vacuum ultraviolet light irradiation to yield zero valence metal masses. An eximer lamp or an eximer laser may preferably used as the light source, and the wavelength range is from about 170 nm to 260 nm.

[0049] The irradiation subject and the light source are placed under ambient pressure or vacuum (120 Pa to 1,200 Pa), and organic materials including the micro phase separation membrane are removed using vacuum ultraviolet light irradiation. The organic materials or their decomposition products are vaporized, and evacuation at a rate of from 1 liter/min to 100 liters/min is preferred to maintain the degree of vacuum at about 120 Pa to 1,200 Pa).

[0050] Metal ions are similarly reduced when ultraviolet light radiation with wavelengths of from 200 nm to 400 nm is used, but an etching treatment to remove organic materials needs to be conducted separately. As such an etching treatment, the micro phase separation membrane is, for example, eluted following the ultraviolet light irradiation using a solvent such as water, ethanol, toluene, methylene chloride and the like. A high pressure mercury lamp, xenon lamp and the like may be used as the ultraviolet light source in these cases.

(2) Electron Beam Irradiation not being part of the present invention.

[0051] When a sample (a micro phase separation membrane containing metal ions) is placed under ambient pressure (in an inert gas atmosphere such as nitrogen and the like) or in vacuum and irradiated using electron beam at a low acceleration voltage (75 kV or lower), the metal ions are reduced and the membrane is decomposed due to the interaction (electrolytic dissociation•excitation) between metal ions and membrane components to form dots and pillars on the substrate.

(3) Plasma Treatment not being part of the present invention

[0052] A reaction gas such as O₂, N₂, H₂, CF₄, CHF₃, BCl₃, SiCl₃, SF₆ and the like is introduced under reduced pressure into a reaction chamber in which a sample (a micro phase separation membrane containing metal ions) is placed, and a plasma is generated by applying a high frequency electrical field or the like. The reaction gas is collisions of electrons in the plasma, and the radicals, ions and the like formed reach the sample and are diffused inside the sample to reduce the metal ions to form dots and pillars. Simultaneously, the membrane is dissipated in the form of highly volatile reaction products.

[0053] As the plasma source, parallel planar type, two cycle parallel planar type, magnetron type, electron cyclotron plasma (ECR), induction coupled plasma (ICP), surface wave plasma (SWP) and the like may be cited.

[0054] In addition, reactive ion beam (RIB) and focused ion beam (FIB) methods that aggressively utilize activated ions are also useful.

(4) Chemical Reduction Process not being part of the present invention

[0055] A micro phase separation membrane containing metal ions is impregnated at room temperature in a solution of a reducing agent such as sodium borohydrate (NaBH₄), sodium thiocyanurate (NaSCN) and the like. The reducing agent penetrates into the hydrophilic sections of the membrane, and the metal ions inside the hydrophilic sections are reduced. The membrane is subsequently gradually eluted, and dots and pillars are formed on the substrate.

(5) Electrochemical Reduction Process not being part of the present invention

[0056] After a micro phase separation membrane containing metal ions is formed on an anode, an electrical field is applied in an electrolyte solution obtained by dissolving an electroconductive salt that does not cause electrical decomposition at the applied potential of the salts of alkali metals or alkaline earth metals, tetralkyl ammonium salts and the like. The metal ions inside the micro phase separation membrane form dots and pillars on the reduced electrode.

[0057] In addition, the metal ion precursors sometimes contain inorganic anions such as F⁻, Cl⁻, Br⁻, I⁻, OH⁻, CN⁻, NO₃⁻, N₃⁻, SCN⁻, SO₄²⁻, CO₃²⁻, acetate, trifluoroacetate, acetylacetonate, carbonate, perchlorate, nitrate, sulfate, oxide and the like or inorganic cations such as K⁺, Na⁺, Rb⁺, Cs⁺, NH₄⁺ and the like as described above. After they are converted into highly volatile compounds they are thought to be removed in the organic material removal step and naturally dissipated or removed upon dissolving in the reducing solution. However, they may be subsequently removed using a washing process if they remain on the substrate.

[0058] As shown in Figure 1, only zero valence metal remains, as a result, in the location corresponding to the cylinder segments of the micro phase separation membrane on a substrate. The shape of the remaining metal masses is the same shape as the hydrophilic cylinders in the micro phase separation membrane or the shape obtained through shrinkage or deformation upon decomposition and removal of the micro phase separation membrane. In addition, the shape may be different depending upon the concentration of the metal present in the cylinder section. The shape is thought to be cylinder-shaped when the concentration is high, and the shape is thought to be sphere-shaped when the concentration is low.

[0059] As a result, a substrate with numerous fine metal masses lined neatly on the surface at 10 nm to 50 nm intervals wherein the masses have shapes that are approximately pillar-shaped (i.e. pillars), approximately dot-shaped (i.e. dots) or a combination of these with diameters of from 3 nm to 15 nm and the pillars are oriented in a direction substantially perpendicular to the surface can be obtained.

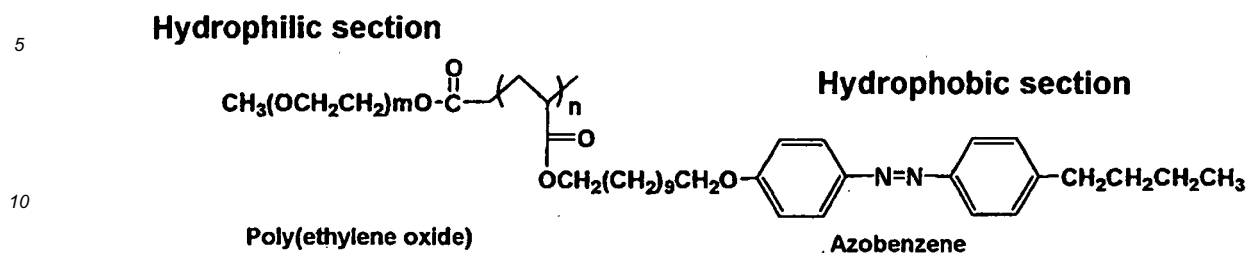
[0060] The present invention is explained below using examples, but the examples are not provided to limit the present invention.

Production Example 1

[0061] Poly(ethylene oxide) methyl ether (molecular weight 5,000) with a degree of polymerization (m) of one hundred and fourteen was used as the hydrophilic polymer chain and was allowed to undergo block copolymerization with benzene methacrylate containing azo groups having hydrophobic and liquid crystalline properties to synthesize an amphiphilic block copolymer with a degree of polymerization (n) of fifty-one. The synthesis was conducted using an atom transfer radical polymerization method and a copper complex as the catalyst.

[0062] The amphiphilic block copolymer obtained is represented by the general formula (Chemical Formula 1) shown below.

[Chemical Formula 1]



15 **[0063]** The number average molecular weight of the copolymer obtained was 30,000, Mw/Mn = 1.23, the polymethacrylate (MA) content was 83% by weight and the melting point was 120°C. (Henceforth referred to as "PEO114-b-PMA(Az)51".)

Production Example 2

20 **[0064]** Poly(ethylene oxide) methyl ether (molecular weight 12,000) with a degree of polymerization (m) of two hundred and seventy-two was used as the hydrophilic polymer chain in the manner described in Production Example 1, and was allowed to undergo block copolymerization with benzene methacrylate containing azo groups having hydrophobic and liquid crystalline properties to synthesize an amphiphilic block copolymer with a degree of polymerization (n) of seventy-four. The synthesis was conducted using an atom transfer radical polymerization method and a copper complex as the catalyst. The number average molecular weight of the copolymer obtained was 48,400, Mw/Mn = 1.10, the polymethacrylate (MA) content was 75% by weight and the melting point was about 120°C. (Henceforth referred to as "PEO272-b-PMA(Az)74".)

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Production Example 3

30 **[0065]** MA (11Az4) 7.38 g (15 mmoles), PEO₂₇₂-BMP 1.716 g (0.143 mmole) as the polymerization initiator and fluorinated resin agitation pieces were added to a 100 ml pear shaped flask with a ground glass neck. The flask along with Cu(I)Cl 63.7 mg (0.643 mmole) in a sample bottle were subjected to three repetitions of nitrogen gas flushing and gas removal at the entrance to a vacuum glove box before they were placed in a glove box. Inside the box, 50 ml of degassed anisole ([M₀]=0.25M) was weighed. A micro syringe was used to add 174.7 μl (148.0 mg, 0.643 mmole) of degassed HMTETA to the Cu(I)Cl in the sample bottle. A pale green copper complex was formed by adding and dissolving a small amount of the weighed anisole. The copper complex solution and the remaining anisole were added to the pear shaped flask, vacuum grease was applied to a ground glass stopper and the ground glass stopper was used to seal the flask. The pear shaped flask was removed from the glove box, and polymerization was allowed to occur for several tens of hours at 80°C.

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40 **[0066]** The polymerization was concluded by opening the pear shaped flask stopper and deactivating the complex by exposing it to air. The solvent was removed using an evaporator. The solids obtained were dissolved in about 50 ml of chloroform, and the copper complex was removed using 200 ml of 1N HCl in four portions and a 500 ml separatory funnel. Next, 5% NaHCO₃ was used four times to neutralize the solution and to confirm the copper complex removal. Lastly, 200 ml of saturated aqueous salt solution (BRAIN) was used once to remove water from the solution. The chloroform solution upon completion of the separation was placed in a 100 ml beaker and was dried by allowing it to stand for an hour after adding magnesium sulfate. The dried chloroform solution was filtered to remove magnesium sulfate. At this point, magnesium sulfate was washed using chloroform until it was white, and the filtrate was treated using an evaporator to remove the solvent. The white solids obtained were washed using boiling hexane to remove unreacted monomer. The washing operation using boiling hexane was assisted by GPC measurements and was repeated until the peak attributed to the monomer disappeared. The solids were dried inside a desiccator to obtain an amphiphilic block copolymer. The number average molecular weight of the copolymer obtained was 62,200, Mw/Mn = 1.10, the polymethacrylate (MA) content was 75% by weight and the melting point was about 120°C. (Henceforth referred to as "PEO272-b-PMA(Az)102".)

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Production Example 4

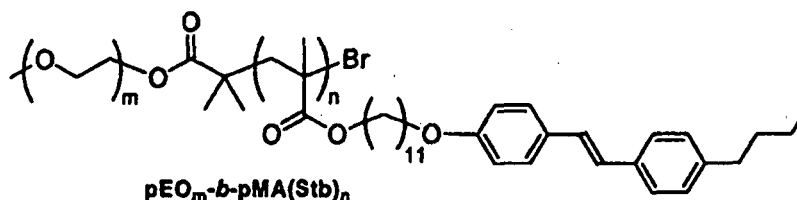
55 **[0067]** MA (11Stb4) 491 mg (10 mmoles), PEO₁₁₄-BMP 714.3 mg (0.143 mmole) as the polymerization initiator and

fluorinated resin agitation pieces were added to a 100 ml pear shaped flask with a ground glass neck. The flask along with Cu(I)Cl 84.9 mg (0.857 mmole) in a sample bottle were subjected to three repetitions of nitrogen gas flushing and gas removal at the entrance to a vacuum glove box before they were placed in a glove box. Inside the box, 40 ml of degassed anisole ($[M_0]=0.25M$) was weighed. A micro syringe was used to add 233 μ l (197 mg, 0.857 mmole) of de-gassed HMTETA to the Cu(I)Cl in the sample bottle. A pale green copper complex was formed by adding and dissolving a small amount of the weighed anisole. The copper complex solution and the remaining anisole were added to the pear shaped flask, vacuum grease was applied to a ground glass stopper and the ground glass stopper was used to seal the flask. The pear shaped flask was removed from the glove box, and polymerization was allowed to occur for sixty hours at 190°C.

[0068] The stopper in the pear shaped flask was opened, and the polymerization was concluded by exposing the complex to air to deactivate the complex. MA(Stb) was slightly soluble in the solvent at ambient temperature, and the solution obtained was warmed to about 40°C. A neutral alumina column ($\Phi=30$ mm, L=100 mm, developing solvent: THF) was used to remove the copper complex, and the solvent was removed using an evaporator. The white solids obtained were washed using boiling hexane to remove unreacted monomer. The washing operation using boiling hexane was assisted by GPC measurements and was repeated until the peak attributed to the monomer disappeared. The solids were dried in a desiccator to obtain an amphiphilic block copolymer (henceforth referred to as "PEO₁₁₄-b-PMA(Stb)₅₂"). [2.36 g, yield 70%, Mn=30,500, Mw/Mn=1.26 (based on GPC), monomer conversion ratio (¹HNMR based) = 54%]

[0069] The amphiphilic block copolymer obtained is represented by the general formula (Chemical Formula 2) shown below.

[Chemical Formula 2]



Example 1 not being part of the present invention

[0070] In the example, gold ion ($HAuCl_4$) was used to form pillars (or dots) comprising zero valence gold on a silicon substrate.

[0071] Chloroauric acid (Wako Junyaku Kogyo) was added to a 3% by weight tetrahydrofuran solution (50 mg, 1,820 μ l) of the amphiphilic block copolymer [PEO₁₁₄-b-PMA(Az)₅₁] obtained in Production Example 1 to prepare a mixed solution. The gold ion in this case was adjusted so that EO/Au = 120 based on the ethylene oxide (EO) constituting the block copolymer.

[0072] The mixed solution prepared as described above was filtered using a syringe filter (Whatman Co., opening 0.45 μ m), and 100 μ l was dropped on a 1 cm x 2 cm silicon substrate. A spin coater (MIKADA Corp., 1H-D7) was used at a rotation rate of 2,000 rpm for ten seconds to apply the cast coating.

[0073] The silicon substrate cast coated with the block copolymer was annealed for twenty-four hours at 120°C under reduced pressure (6.7×10^{-2} Pa) to form a nano phase separation structure. Upon annealing, the nano phase separation membrane was examined using an atomic force microscope (Seiko Instrument Co., SPA400). The results are shown in Figure 3.

[0074] The nano phase separation membrane that had been subjected to gold ion interaction was irradiated for sixty minutes at room temperature under reduced pressure (1.2×10^3 Pa) with 172 nm vacuum ultraviolet light from an irradiation distance of 2 mm using a vacuum ultraviolet light irradiation device (Ushio Denki Co., UEM20-172). The membrane was washed using methylene chloride (Wako Junyaku Kogyo) after the irradiation followed by super pure water (18.2 M Ω ·cm). The substrate IR spectra before and after the vacuum ultraviolet light irradiation obtained using an FT-IR spectroscopy (Nippon Bunko K.K., FT/IR-5300) were examined. The results are shown in Figure 4.

[0075] After the membrane was decomposed and removed, the nano gold dots formed on the substrate were examined using an atomic force microscope (AFM) (Seiko Instruments K.K., SPA400). The results are shown in Figure 5. Highly ordered nano Au dots reflecting PEO₁₁₄-b-PMA(Az)₅₁ structures 10 nm in diameter and 4 nm high located at about 30 nm cycles were obtained.

[0076] Au4f electron peaks were observed at 84 eV in an X-ray photoelectron spectroscopy (XPS) evaluation of the substrate surface.

[0077] The data indicated that the ozone generated by the vacuum ultraviolet light irradiation decomposed and removed

the organic material on the substrate and also reduced the gold ions to metallic gold (zero valence).

Example 2 not being part of the present invention

5 **[0078]** In the example, silver ion (AgNO_3) was used to form pillars (or dots) comprising zero valence silver on a silicon substrate.

[0079] The PEO114-b-PMA(Az)51 obtained in Production Example 1 and AgNO_3 (Wako Junyaku Kogyo) were mixed in toluene to achieve solutions with EO/ Ag^+ molar ratios of twenty and one hundred and twenty. The 2 w/w % toluene solutions were spin coated on substrates to prepare membranes about 100 nm thick. The membranes were subjected to a heat treatment (annealing) and were subsequently irradiated for thirty minutes using a vacuum ultraviolet light. The results are shown in Figure 6.

10 **[0080]** Mock hexagonal nano particles as shown in (a) were observed when the EO/ Ag^+ molar ratio was twenty. The particle size was 10 nm, and the height was 3 nm. The particles size was 6 nm and the height was 2 nm as seen in (b) when the EO/ Ag^+ molar ratio was as low as one hundred and twenty. The observation indicated that the dot size was optionally controlled by changing the concentration.

15 **[0081]** The results of inspecting the properties of Ag particles formed on quartz glass were shown in Figure 7. The peak (A) attributed to PEO114-b-PMA(Az)51 disappeared as indicated in (a) after vacuum ultraviolet light irradiation, and Ag particle plasmon absorption was observed at 450 nm (B).

20 **[0082]** Next XPS evaluation was conducted to confirm that the Ag was reduced to zero valence Ag particles (b). The bonding energy increased somewhat upon irradiation (B) when compared to the sample (A) prior to the vacuum ultraviolet light irradiation. When the sample was subjected to a heat treatment at 600°C in air to oxidize it, the bonding energy declined to the original value (C). The observation indicated that zero valence Ag was formed due to Ag^+ reduction through vacuum ultraviolet light irradiation.

25 Example 3

[0083] In the example, silver ion (CAgNO_3) was used to form pillars (or dots) comprising zero valence silver on a silicon substrate.

30 **[0084]** The PEO114-b-PMA(Az)51 obtained in Production Example 1 was dissolved in toluene to form a 2 w/w % (weight ratio) (molar ratio 6.6×10^{-5}) solution, and the solution was spin coated on quartz or Si (100) to prepare membranes about 100 nm thick. The membrane was subjected to a heat treatment (annealing) for twenty-four hours at 140°C under vacuum (25 Pa). An aqueous 1.0 M AgNO_3 solution was added dropwise onto the membrane and was left standing for thirty minutes at room temperature. The membrane was subsequently washed using distilled water and dried under vacuum (100Pa). Samples obtained in the manner described above were irradiated for thirty minutes using 172 nm vacuum ultraviolet light under reduced pressure (50Pa).

35 **[0085]** The results obtained with AMF when a silicon substrate was used are shown in Figure 8. (a) is the membrane prior to the treatment, and (b) is the membrane treated with Ag^+ . (c) is the topological image of the membrane after vacuum ultraviolet light irradiation, and (d) is the height image. (a) indicated that the block polymer had a cycle of 20 nm and a PEO cylinder diameter of 10 nm. (b) indicated that Ag^+ was selectively ligand bonded to the PEO segments, and its construction was not changed by the treatment. After the vacuum ultraviolet light irradiation, PEO114-b-PMA(Az)51 was completely removed (confirmed by infrared spectroscopy), and highly ordered AgO nano dots reflecting a PEO114-b-PMA(Az)51 structure with a cycle of about 18 nm, particle diameter of 8 nm and height of 3 nm were obtained.

45 Example 4 not being part of the present invention

[0086] In the example, palladium ion (palladium acetate) was used to form pillars (or dots) comprising zero valence palladium.

50 **[0087]** The amphiphilic block copolymer, PEO272-b-PMA(Az)74, obtained in Production Example 2 was used to prepare a 3.0% by weight toluene solution, and palladium acetate, $\text{Pd}(\text{acac})_2$, was used to prepare a 0.5% by weight toluene solution. The solutions were combined and the combined solution was agitated for three hours at room temperature. The palladium ion was adjusted so that EO/Pd=16.36, 5.54, 3.27 based on the ethylene oxide (EO) constituting the block copolymer at this point.

55 **[0088]** Next a silicon substrate (Fuji Fine Technology Co., azimuth one hundred facets, thickness $525 \mu\text{m} \pm 25 \mu\text{m}$), the surface of which was made hydrophilic, was separately prepared by irradiating the substrate for ninety minutes at an irradiation distance of 2 mm using 172 nm vacuum ultraviolet light from a vacuum ultraviolet light irradiation device (Ushio Denki K.K., UEM20-172) under reduced pressure (170 Pa).

[0089] The mixed solution prepared as described above was filtered using a syringe filter (Whatman Co., opening $0.45 \mu\text{m}$), and 100 μl was subsequently added dropwise to a hydrophilic silicon substrate 1 cm x 2 cm. A spin coater

(Mikasa K.K., 1H-D7) was used at a rotary speed of 2,000 rpm for ten seconds to create a coating. The silicon substrate coated with the block copolymer was annealed for twelve hours at 140°C under reduced pressure (6.7×10^{-2} Pa). The nano phase separation membrane obtained upon annealing was examined using an atomic force microscope (Seiko Instruments K.K., SPA400). The results are shown in Figure 9. The data revealed that a nano phase separation structure was formed.

[0090] The nano phase separation membrane obtained by the interaction with $\text{Pd}(\text{acac})_2$ was irradiated for ninety minutes at an irradiation distance of 2 mm using a 172 nm vacuum ultraviolet light from a vacuum ultraviolet light irradiation device (Ushio Denki K.K., UEM20-172) under reduced pressure (170 Pa) at room temperature. After the membrane was decomposed and removed, the substrate was washed with methylene chloride (Wako Junyaku Kogyo) and subsequently with super pure water ($18.2 \text{ M}\Omega\cdot\text{cm}$). The palladium dots formed on the substrate were examined using an atomic force microscope (Seiko Instruments K.K., SPA400). The results are shown in Figure 10. Highly ordered Pd nano dots reflecting a PEO272-b-PMA(Az)74 structure with a cycle of about 40 nm, particle diameter of 17 nm and height of 3 nm were obtained. The distance between centers of the pillars (or dots) formed was 35 nm.

[0091] According to the X-ray photoelectron spectroscopic measurements of the substrate surface, a Pd4f electron peak was observed at 335 eV. The data obtained revealed that the organic materials on the substrate were decomposed and removed by the ozone generated with the vacuum ultraviolet light irradiation and the palladium ion was reduced to metallic palladium (zero valence).

Example 5 not being part of the present invention

[0092] A mixed solution was prepared by adding chloroauric acid (HAuCl_4) to a 3% by weight toluene solution of the amphiphilic block copolymer, PEO272-b-PMA(Az)102, obtained in Production Example 3 so that $\text{EO}/\text{Au} = 200$. In addition, a silicon substrate (Fuji Fine Technology Co., azimuth one hundred facets, thickness $525 \mu\text{m} \pm 25 \mu\text{m}$) was separately prepared. The mixed solution prepared as described above was filtered using a syringe filter (Whatman Co., opening $0.45 \mu\text{m}$), and 100 μl was subsequently added dropwise to a hydrophilic silicon substrate 1 cm x 2 cm. A spin coater (Mikasa K.K., 1H-D7) was used at a rotary speed of 2,000 rpm for ten seconds to create a coating. The silicon substrate coated with the block copolymer was annealed for twenty-four hours at 140°C to prepare a nano phase separation structure.

[0093] Next the nano phase separation membrane obtained by allowing the gold ion to interact was impregnated for several tens of seconds in an aqueous 0.1M sodium borohydride solution (NaBH_4) to reduce the gold ion and the remove the membrane.

[0094] The nano phase separation membrane obtained was examined using an atomic force microscope (Seiko Instruments K.K., SPA400). The results are shown in Figure 11. (a) indicates the nano phase separation membrane doped with gold ion, and (b) indicates the membrane after a sixty second exposure to the aqueous sodium borohydride solution.

[0095] Furthermore, the results obtained upon exposing the membrane for sixty seconds to the aqueous sodium borohydride solution and examining the product using a transmission electron microscope (Hitachi K.K., Hitachi H-800) are shown in Figure 12. The metallic gold, which appears in black, spreads hexagonally, and the distance was in good agreement with the cylinder cycle $d_{\text{PEO}} = 38 \text{ nm}$ of the nano phase separation membrane PEO₂₇₂-b-PMA(Az)₁₀₂ used as the template indicating that the gold ion incorporated into the PEO cylinder interior was reduced to metallic gold inside the cylinders.

[0096] In addition, the residual block copolymer was completely removed using an excimer VUV/O₃ washing device (Ushio Denki K.K., UEM20-172). The atomic force microscope image after the membrane was removed is shown in Figure 11c. Highly ordered Au nano dots reflecting a PEO272-b-PMA(Az)102 structure of the amphiphilic block copolymer was obtained.

Example 6 not being part of the present invention

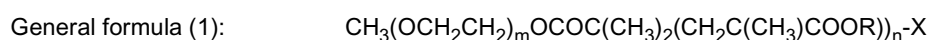
[0097] A mixed solution was prepared by adding chloroauric acid (HAuCl_4) to a 2% by weight toluene solution of the amphiphilic block copolymer, PEO₁₁₄-b-PMA(Stb)₅₂, obtained in Production Example 3 so that $\text{EO}/\text{Au} = 200$. The mixed solution prepared as described above was filtered using a syringe filter (Whatman Co., opening $0.45 \mu\text{m}$), and 100 μl was subsequently added dropwise to a hydrophilic silicon substrate 1 cm x 2 cm. A spin coater (Mikasa K.K., 1H-D7) was used at a rotary speed of 2,000 rpm for ten seconds to create a coating. The silicon substrate coated with the block copolymer was annealed for two hours at 190°C under reduced pressure to prepare a nano phase separation structure. The nano phase separation membrane obtained after annealing was examined using an atomic force microscope (Seiko Instruments K.K., SPA400). The results are shown in Figure 13a.

[0098] Next the nano phase separation membrane that had been subjected to gold ion interaction was irradiated for sixty minutes at room temperature under reduced pressure ($1.2 \times 10^3 \text{ Pa}$) with 172 nm vacuum ultraviolet light from an

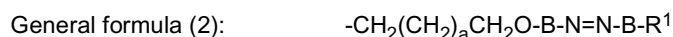
irradiation distance of 20 mm using a vacuum ultraviolet light irradiation device (Ushio Denki Co., UEM20-172). The ozone generated during irradiation decomposed and removed the organic materials on the substrate, and the gold ion was also reduced to metallic gold. The IR spectra of the substrate before and after the vacuum ultraviolet light irradiation obtained using an FT-IR spectroscope (Nippon Bunko K.K., FT/IR-5300) were examined. The results are shown in Figure 14. In addition, the results of X-ray photo electron spectroscopic measurements of the substrate surface are shown in Figure 15. The Au4f electron peak was observed at 84 eV. The data revealed that the organic materials on the substrate were decomposed and removed by the ozone generated, and the metal ion was reduced to metallic gold (zero valence). [0099] After the membrane was decomposed and removed, the nano gold dots formed on the substrate were examined using an atomic force microscope (AFM) (Seiko Instruments K.K., SPA400). The results are shown in Figure 13b. Highly orderly nano Au dots reflecting the PEO₁₁₄-b-PMA(Stb)₅₂ structures of the amphiphilic block copolymer were obtained.

Claims

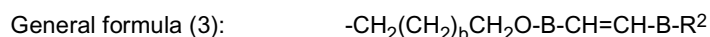
1. A method for preparing a substrate with micrometallic masses aligned on its surface, wherein the substrate contains numerous zero valence fine metal masses aligned on the surface at intervals of 10 nm to 50 nm wherein the fine metal masses are pillar-shaped or dot-shaped or a combination thereof and have diameters of 3 nm to 15 nm wherein the pillar-shaped masses are oriented in a direction substantially perpendicular to the surface, wherein the method comprises the steps of (a) preparing a micro phase separation membrane of an amphiphilic block copolymer on a substrate wherein the membrane has a thickness of 50 nm to 1 μm and then exposing the membrane to a solution containing metal ions, (b) drying the membrane to form a micro phase separation membrane with the metal ions localized in its hydrophilic segment, and (c) irradiating vacuum ultraviolet light on the micro phase separation membrane under vacuum to reduce the metal ions and, as a result of the irradiation, organic materials including those in the micro phase separation membrane are removed, leaving the fine metal masses on the surface.
2. The method of claim 1 wherein the metal is a copper group metal or a platinum group transition metal.
3. The method of claim 1 or 2 wherein the amphiphilic block copolymer is represented by the general formula (1) below:



wherein m and n, may be identical or different from each other, are integers 5 to 500 independently, R is a substituent represented by the general formula (2) or (3) shown below, X represents a hydrogen atom or halogen atom;



wherein a is an integer 0 to 20, R¹ is a hydrogen atom or an alkyl group containing 1 to 22 carbon atoms and B represents a p-phenylene group;



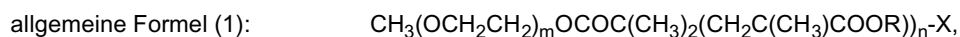
wherein b is an integer 0 to 20, R² is a hydrogen atom or an alkyl group containing 1 to 22 carbon atoms and B represents a p-phenylene group.

Patentansprüche

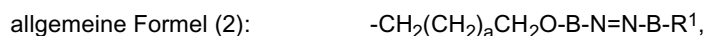
1. Verfahren zur Herstellung eines Substrats mit mikrometallischen Massen, die auf dessen Oberfläche angeordnet sind, wobei das Substrat eine Vielzahl von feinen Nullvalenz-Metallmassen beinhaltet, welche auf der Oberfläche mit Abständen von 10 nm bis 50 nm angeordnet sind, wobei die feinen Metallmassen säulenförmig oder punktförmig oder eine Kombination von beiden sind und Durchmesser von 3 nm bis 15 nm besitzen, wobei die säulenförmigen Massen in einer zur Oberfläche im Wesentlichen senkrechten Richtung orientiert sind, wobei das Verfahren die Schritte umfasst: (a) Präparieren einer Mikrophasentrennmembran eines amphiphilen Block-Copolymers auf einem Substrat, wobei die Membran eine Dicke von 50nm bis 1 μm hat, und dann das Aussetzen der Membran einer Metallionen enthaltenden Lösung, (b) Trocknen der Membran, um eine Mikrophasentrennmembran auszubilden, wobei die Metallionen auf ihrem hydrophilen Segment lokalisiert sind, und (c) Einstrahlen von ultravioletter Licht bei Vakuum auf die Mikrophasentrennmembran unter Vakuum, um die Metallionen zu reduzieren, und wobei als ein Ergebnis der Bestrahlung organische Materialien, eingeschlossen jener in der Mikrophasentrennmembranen,

entfernt werden, die feinen Metallmassen auf der Oberfläche zurücklassend.

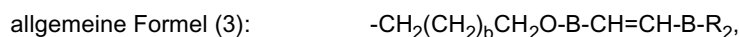
2. Verfahren gemäß Anspruch 1, wobei das Metall ein Metall der Kupfergruppe oder ein Übergangsmetall der Platin-
gruppe ist.
3. Verfahren gemäß Anspruch 1 oder 2, wobei das amphiphile Block-Copolymer durch die folgende untenstehende
allgemeine Formel (1) wiedergegeben ist:



wobei m und n,
welche identisch oder verschieden sein können, unabhängig voneinander ganze Zahlen von 5 bis 500 sind, R ein
durch die unten gezeigte allgemeine Formel (2) oder (3) dargestellter Substituent ist, X ein Wasserstoffatom oder
ein Halogenatom darstellt;



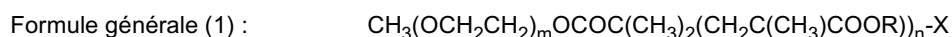
wobei a eine ganze Zahl von 0 bis 20 ist, R¹ ein Wasserstoffatom oder eine 1 bis 22 Kohlenstoffatome enthaltende
Alkylgruppe ist, und B eine p-Phenylengruppe darstellt;



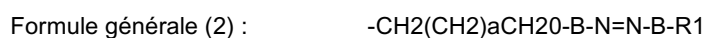
wobei b eine ganze Zahl von 0 bis 20 ist, R² ein Wasserstoffatom oder eine 1 bis 22 Kohlenstoffatome enthaltende
Alkylgruppe ist, und B eine p-Phenylengruppe darstellt.

Revendications

1. Procédé de préparation d'un substrat avec des masses micro-métalliques alignées sur sa surface, dans lequel le
substrat contient de nombreuses masses en métal fin de valence zéro alignées sur la surface à des intervalles de
10 nm à 50 nm, dans lequel les masses en métal fin sont en forme de piliers ou de points ou d'une combinaison
de ceux-ci, et ont des diamètres de 3 nm à 15 nm, dans lequel les masses en forme de piliers sont orientées dans
une direction sensiblement perpendiculaire à la surface, dans lequel le procédé comprend les étapes (a) de prépa-
ration d'une membrane de séparation de micro phase d'un copolymère à bloc amphiphile sur un substrat, dans
laquelle la membrane a une épaisseur de 50 nm à 1 µm, et ensuite d'exposition de la membrane à une solution
contenant des ions métalliques, (b) de séchage de la membrane afin de former une membrane de séparation de
micro phase avec les ions métalliques localisés sur son segment hydrophile, et (c) d'irradiation à la lumière ultra-
violette sous vide de la membrane de séparation de micro phase sous vide pour réduire les ions métalliques, et,
en conséquence de l'irradiation, des produits organiques incluant ceux dans la membrane de séparation de micro
phase sont retirés, laissant les masses de métal fin sur la surface.
2. Procédé selon la revendication 1, dans lequel le métal est un métal du groupe du cuivre ou un métal de transition
du groupe du platine.
3. Procédé selon les revendications 1 ou 2, dans lequel le copolymère à bloc amphiphile est représenté par la formule
générale (1) ci-dessous :



où m et n peuvent être identiques ou différents l'un de l'autre, sont des nombres entiers de 5 à 500 de manière
indépendante, R est un substituant représenté par la formule générale (2) ou (3) exposée ci-dessous, X représente
un atome d'hydrogène ou un atome d'halogène ;



où a est un nombre entier de 0 à 20, R¹ est un atome d'hydrogène ou un groupe alkyle contenant 1 à 22 atomes
de carbone et B représente un groupe p-phénylène ;

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Formule générale (3) : $-\text{CH}_2(\text{CH}_2)_b\text{CH}_2\text{-B-CH=CH-B-R}_2$

où b est un nombre entier de 0 à 20, R² est un atome d'hydrogène ou un groupe alkyle contenant 1 à 22 atomes de carbone et B représente un groupe p-phénylène.

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Figure 1

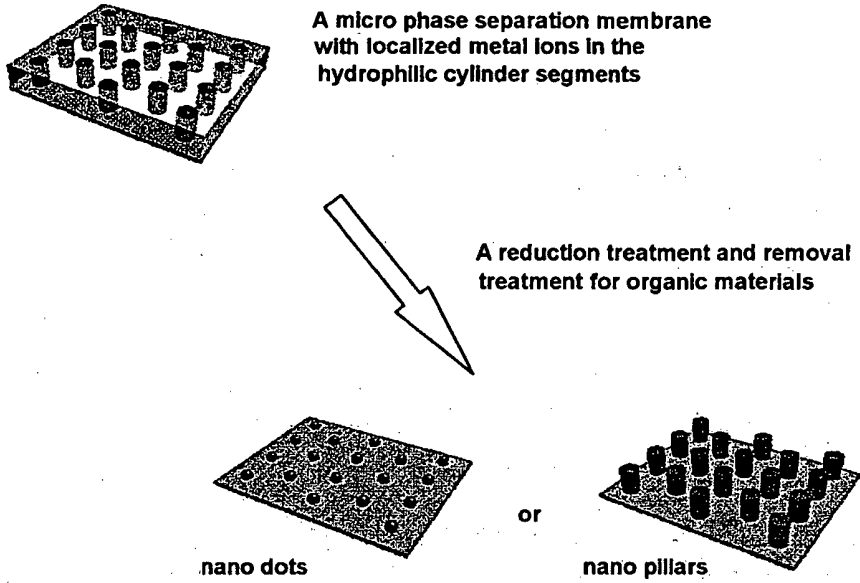


Figure 2

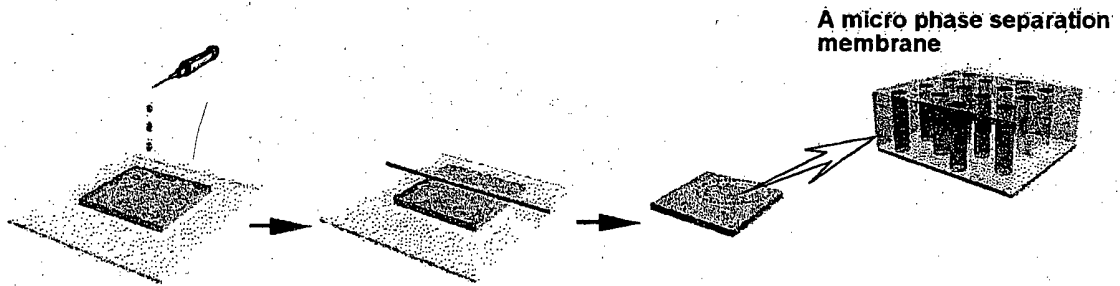


Figure 3

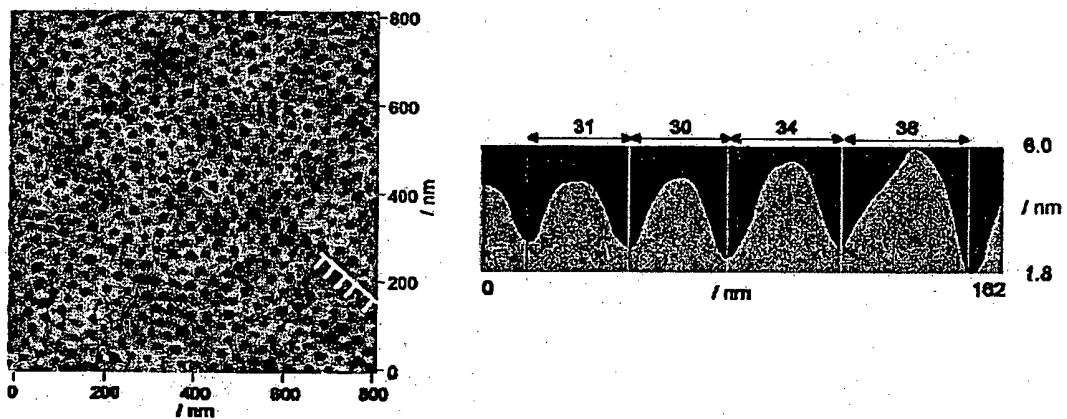


Figure 4

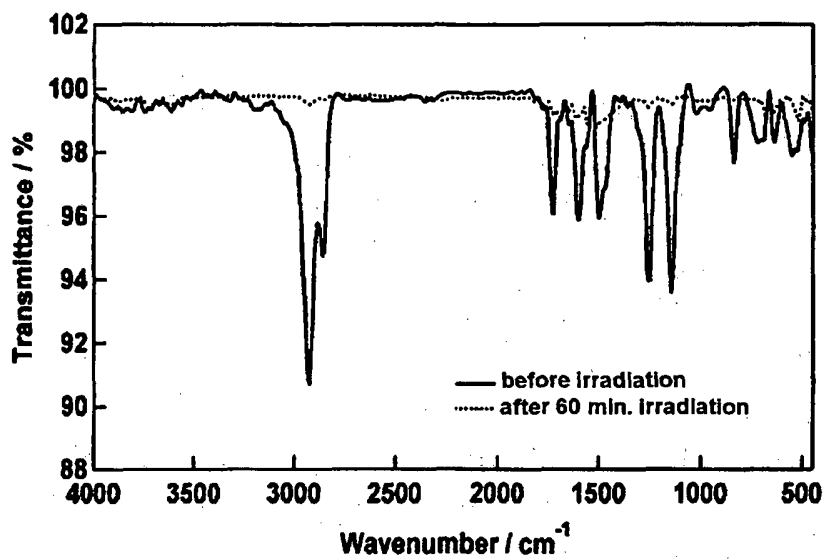


Figure 5

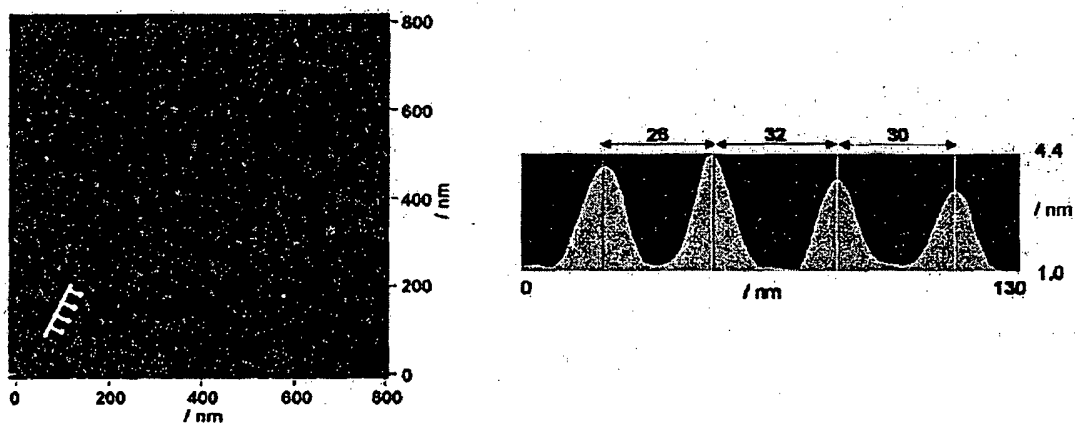


Figure 6

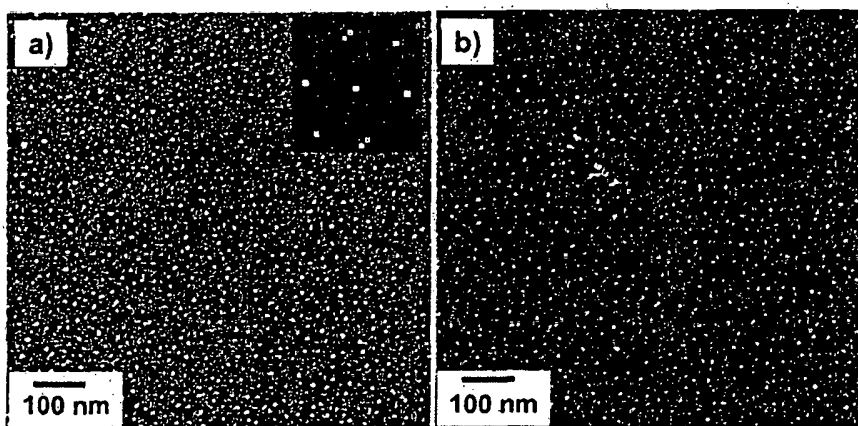


Figure 7

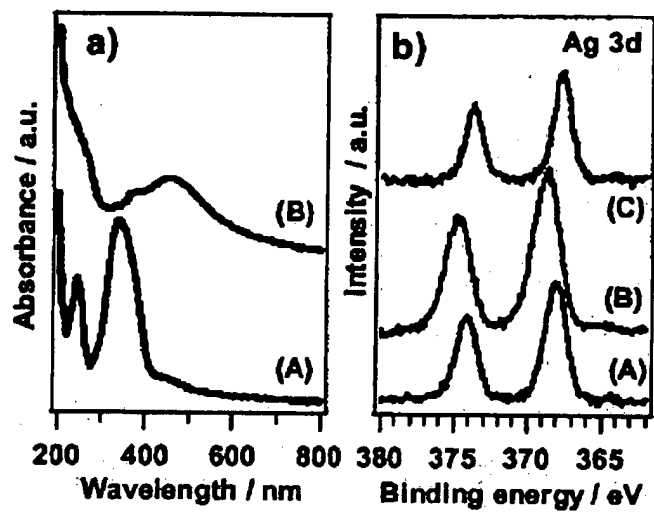


Figure 8

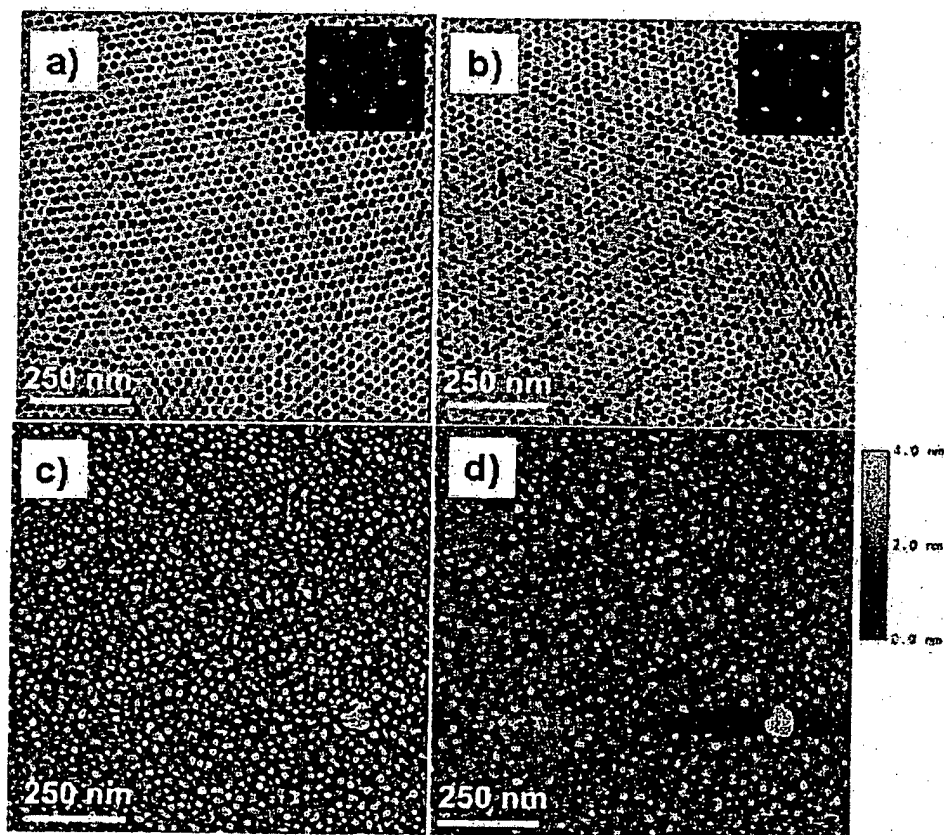


Figure 9

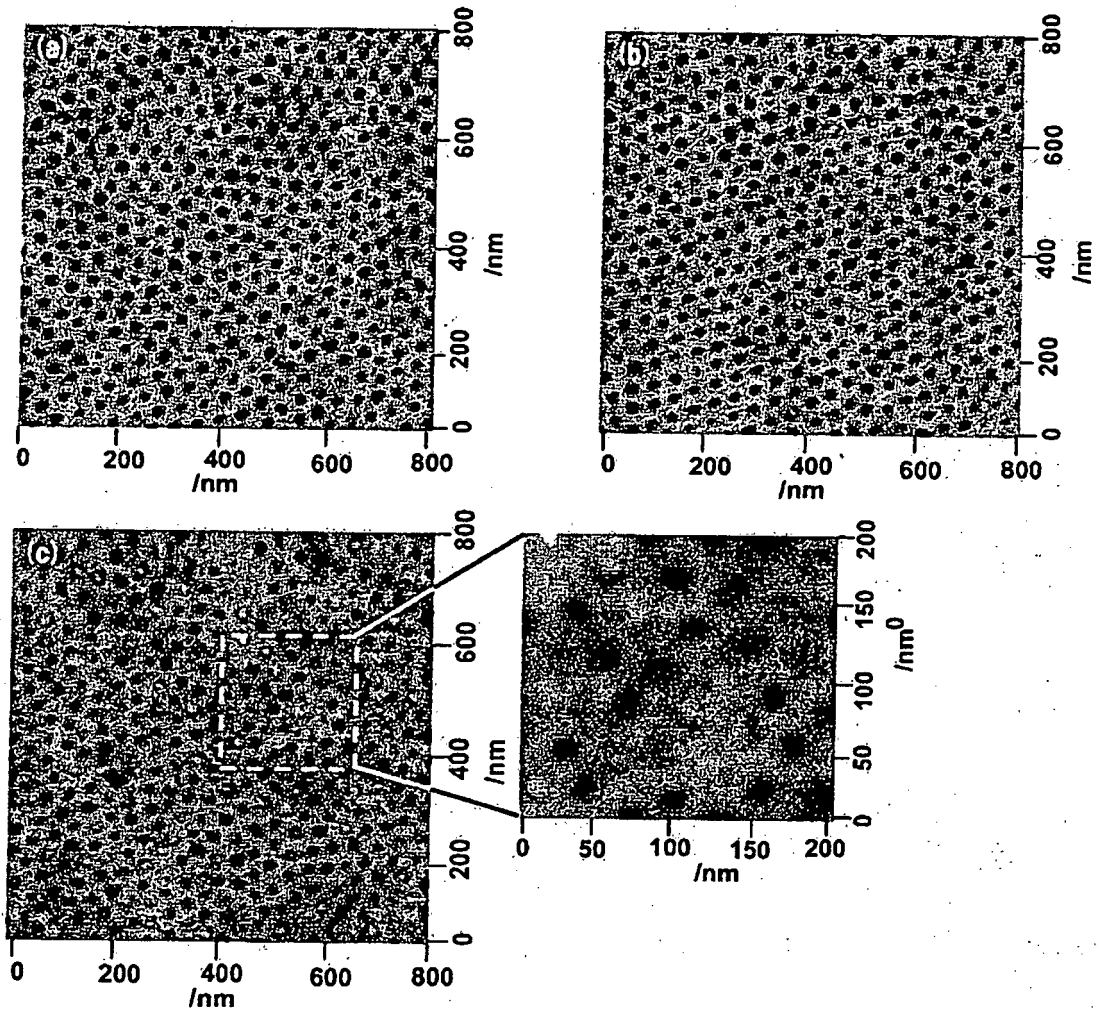


Figure 10

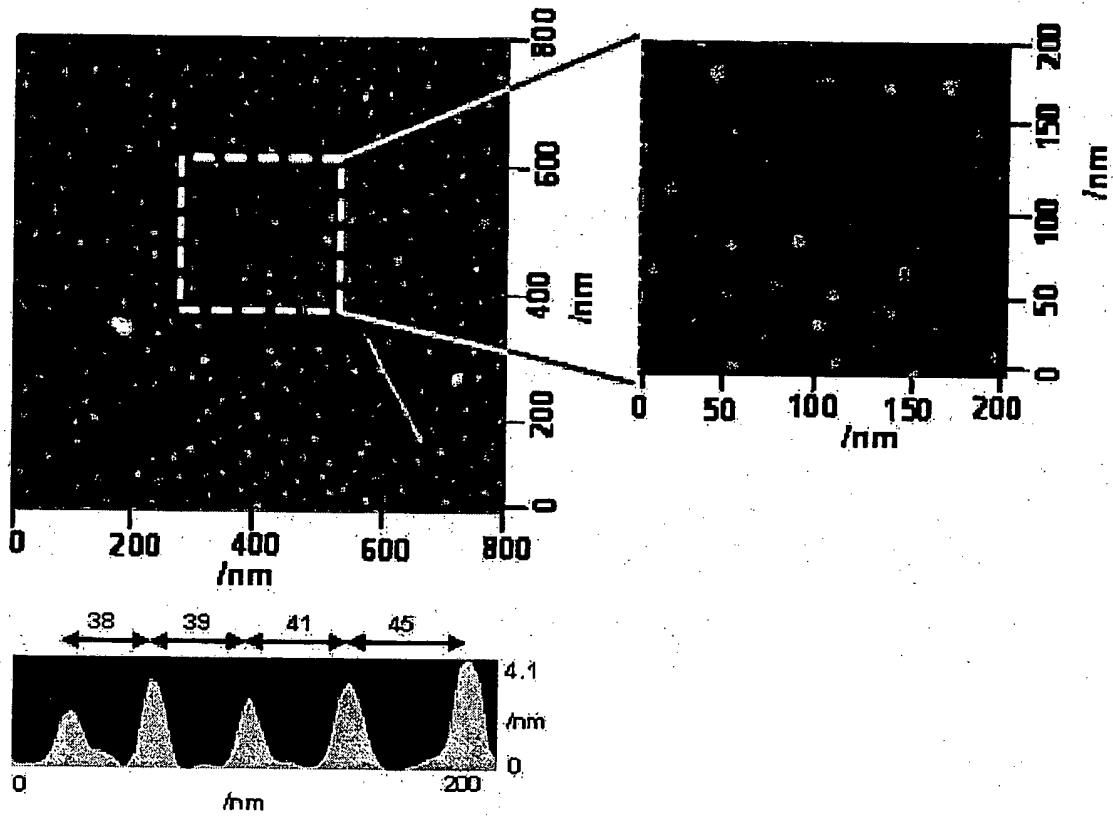


Figure 11

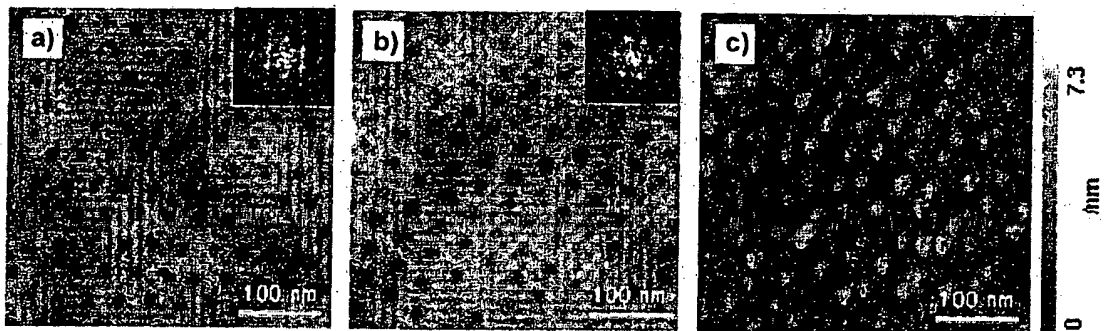


Figure 12

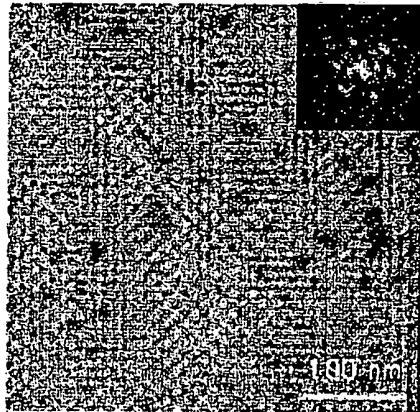


Figure 13

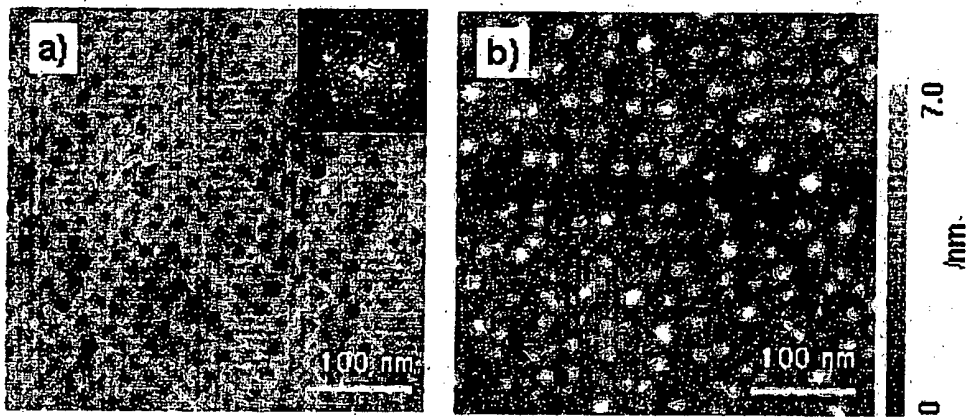


Figure 14

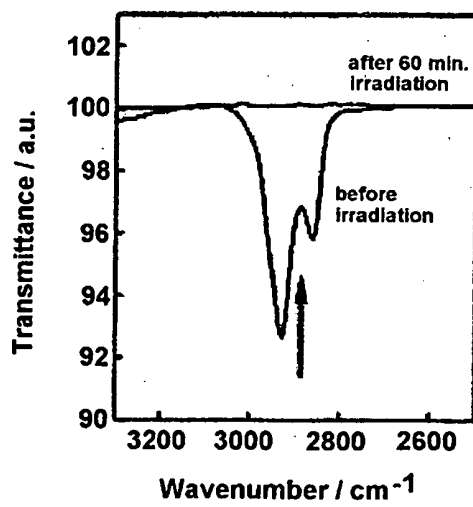
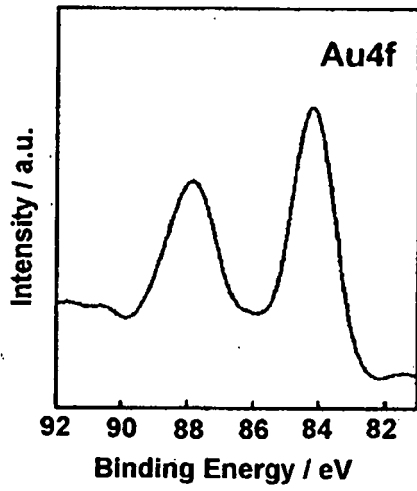


Figure 15



REFERENCES CITED IN THE DESCRIPTION

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