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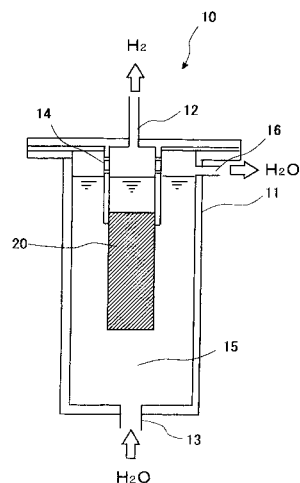
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(54) **HYDROGEN GAS GENERATING MEMBER AND HYDROGEN GAS PRODUCING METHOD THEREFOR**

(57) There is provided a hydrogen gas generating member which safely facilitates the hydrogen gas generation reaction by bringing an Al alloy which is subjected to rolling treatment or powdering treatment into contact with water.

A hydrogen gas generating member 20 includes a texture in which Al is finely dispersed in a metal matrix, where hydrogen gas is generated by bringing the hydrogen gas generating member into contact with water. A fixing member 14 for mounting the hydrogen gas generating member 20 is provided in a hydrogen generating apparatus 10 and is brought into contact with a water 15 that is stored inside. The hydrogen gas generated from the surface is supplied outside through a hydrogen gas collecting pipe 12 and stored in a storage tank (not shown).

Fig. 1



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Description

Technical Field

5 **[0001]** The present invention relates to a hydrogen gas generating member which produces hydrogen gas to be used as a fuel for motors and fuel cells or a raw material for chemical industries using the reaction of a hydroxide forming metal typified by Al with water and its process for producing hydrogen gas.

Background Art

10 **[0002]** Hydrogen gas has been conventionally produced by electrolysis of water ($H_2O \rightarrow H_2 + 1/2O_2$) or steam reforming of hydrocarbon ($CH_4 + H_2O \rightarrow 3H_2 + CO$). However, when the electrolysis method is used, the energy efficiency is low. When the steam reforming process is used, there are problems such as a large burden on the facility and disposal of by-products. Thus, a process for producing hydrogen by reaction of metal and water has been examined.

15 Typical examples of the metal to be used for hydrogen production include Al and an Al alloy. For example, in Patent documents 1, 2, and 4, hydrogen gas is generated and trapped by the steps of immersing Al in Ga with a melting point of 30°C or less or a Ga-In alloy melt, removing a surface oxide film, exposing an active surface, and contact-reacting Al with water ($Al + 3H_2O \rightarrow Al(OH)_3 + 1.5H_2$). Further, a method including the steps of cutting and processing an Al alloy immersed in water and exposing a fresh active surface (Patent document 3), friction and dynamic fracture associated with friction (Patent document 5), thermal shock by rapid heating and quenching (Patent document 6) are known.

20 **[0003]** When the generation of hydrogen gas is dependant on the reaction of the active surface of the Al alloy and water, the hydrogen gas yield is influenced by the size of the active surface. It is said that the bulk of the Al alloy does not contribute to a hydrogen gas generation reaction. Since the size of the bulk is overwhelmingly larger than that of the active surface, the hydrogen gas yield is significantly lower than the consumption of the Al alloy. It may remain within several percent of a theoretical value of 1.31/g·Al which is calculated from a reaction formula of $Al + 3H_2O \rightarrow Al(OH)_3 + 1.5H_2$. Further, it is reported that hydrogen gas is generated by contact-reacting metals having a low melting point, such as Ga, In, Sn, or Zn and the alloyed Al with water (Nonpatent document 1). From the viewpoint that the hydrogen gas generation reaction is occurred by bringing dopants such as Ga, In, Sn, or Zn and the alloyed Al into contact with water, the process is advantageous as compared to a method depending on the reaction of an active surface which is newly formed by cutting and frictioning the surface of the Al alloy with water. However, the alloyed Al alloy contains expensive metals such as Ga and In and thus it is not suitable from a cost standpoint.

25 Further, it is reported that the hydrogen gas generation reaction is induced by bringing alloy powder produced by adding 0 to 30% by weight of Bi to Al by the mechanical alloying process using a ball mill into contact with water (Nonpatent document 2).

30 **[0004]**

Patent document 1: U.S. Patent No. 4358291

Patent document 2: U.S. Patent No. 4745204

Patent document 3: Japanese Patent Application Laid-Open (JP-A) No. 2001-31401

35 Patent document 4: JP-A No. 2003-12301

Patent document 5: JP-A No. 2004-123517

Patent document 6: JP-A No. 2006-45004

Nonpatent document 1: O.V Kravchenko et al., Journal of Alloys and Compounds 397(2005). pp.58-62

40 Nonpatent document 2: M-Q.Fan et al., International Journal of Hydrogen Energy, Corrected Proof, Available online 12 February 2007

Disclosure of the Invention

Problems to be solved in the Invention

50 **[0005]** The fact that the hydrogen gas yield is changed by the water temperature in this case is described in Patent document 6 and Nonpatent document 1. From an economical point of view, it goes without saying that it is desirable to obtain a stable hydrogen yield when the water temperature is room temperature (from 5 to 40°C).

55 **[0006]** The present inventors have investigated and examined various effects caused by alloying of dopant. As the result, they have found that, surprisingly, when the alloyed Al alloy is subjected to a rolling treatment or a powdering treatment, the hydrogen gas generation reaction is proceeded safely and efficiently without expensive Ga.

Thus, an objective of the present invention is to provide a hydrogen gas generating member which safely promotes the hydrogen gas generation reaction by performing the rolling treatment or the powdering treatment and bringing into

contact with water based on the finding.

Further, the objective of the present invention is to provide a process for producing hydrogen gas which includes a step of performing the hydrogen gas generation reaction by bringing into contact with water using the hydrogen gas generating member.

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Means to solve the Problems

[0007] In order to solve the above-described problems, according to the present invention, there is provided a hydrogen gas generating member including a texture in which Al is finely dispersed in a metal matrix, where hydrogen gas is generated by bringing the hydrogen gas generating member into contact with water.

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According to the present invention, the hydrogen gas generating member is separated into two-liquid phases having mutually different compositions in a liquid phase state and it is separated into two-liquid phases in a solid phase state. According to the present invention, the hydrogen gas generating member is an Al-X alloy and X is at least one selected from Sn, Bi, and In.

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According to the present invention, the hydrogen gas generating member is an Al-X alloy and X is selected from Sn: 10.1 to 99.5% (mass percent: hereinafter on the basis of mass percent), Bi: 30.1 to 99.5%, In: 10.1 to 99.5%, Sn+Bi: 20.1 to 99.5%, Sn+In: 10 to 99.5%, Bi+In: 20.1 to 99.5%, and Sn+Bi+In: 20 to 99.5%.

20

According to the hydrogen gas generating member of the present invention, the Al-X alloy further includes Y and Y is at least one selected from Pb: 0.01 to 20%, Zn: 0.01 to 30%, Si: 0.01 to 20%, Cd: 0.01 to 20%, Cu: 0.01 to 20%, Fe: 0.01 to 5%, Ga: 0.01 to 30%, Ge: 0.01 to 30%, Hg: 0.01 to 20%, Mg: 0.01 to 20%, Ni: 0.01 to 5%, S: 0.01 to 5%, and Mn: 0.01 to 5%, the content is within the range of 60% even when a plurality of elements are contained, and the relation of X and Y is defined by $2X > Y$.

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According to the present invention, the hydrogen gas generating member is rolled and processed.

According to the present invention, the hydrogen gas generating member is powdered.

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[0008] A process for producing hydrogen gas includes a step of bringing the hydrogen gas generating member according to any one described above into contact with water to generate hydrogen gas.

Effects of the Invention

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[0009] According to the means to solve the problems, when the hydrogen gas generating member of the present invention is used, the hydrogen gas generation reaction may be produced with inexpensive components. In addition, the hydrogen gas generation reaction may be safely produced by bringing into contact with water.

Further, the process for producing hydrogen gas of the present invention has a simple structure and may allow the hydrogen gas generation reaction to produce safely.

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Brief Description of the Drawings

[0010]

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Fig. 1 is a schematic diagram showing a structure of a hydrogen gas generating apparatus which generates hydrogen gas.

Fig. 2 is an SEM photograph showing the cross section of the hydrogen gas generating member containing Al and Sn.

Fig. 3 is an SEM photograph showing the cross section of the hydrogen gas generating member containing Al, Sn, and In.

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Fig. 4 is an SEM photograph of the cross-sectional texture of the hydrogen gas generating member containing Al, Sn, and In when hydrogen is generated.

Description of the Reference Numerals

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[0011]

10 Hydrogen generating apparatus

11 Apparatus container

12 Hydrogen gas collecting pipe

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13 Water supply pipe

14 Fixed member

15 Water

16 Water drain pipe

20 Hydrogen gas generating Member

Best Mode for carrying out the Invention

5 **[0012]** Hereinafter, the best mode for carrying out the invention will be described based on the drawings. In this regard, it is easy for a so-called person skilled in the art to make other embodiments by using variations and modifications of the present invention in the claims. These modifications and variations are intended to fall within the scope of the claims, but do not constitute a limitation of the claims.

10 The hydrogen gas generating member of the present invention includes the texture in which Al is finely dispersed in the metal matrix, where hydrogen gas is generated by bringing the hydrogen gas generating member into contact with water. In Al in which a clean surface is exposed, it is known that hydrogen is generated while Al is reacted with water to be converted to $\text{Al}(\text{OH})_3$. Similarly, it is known that when a high energy light having a short wavelength such as sunlight is given in the action, water is decomposed to generate hydrogen. The hydrogen gas generating member of the present invention generates hydrogen using a successive hydrogen generation mechanism in which hydrogen is generated
15 while Al which is finely dispersed in the metal matrix is reacted with water to be converted to $\text{Al}(\text{OH})_3$, water permeates into a part where Al disappeared, and Al in the metal matrix reacts with water. When among them, Al is dispersed in the metal matrix, the hydrogen generating capability is not significantly reduced even if the Al surface is oxidized by oxygen. Al is formed into a fine particle size and dispersed in the metal matrix. With reference to the fine texture of Al (hereinafter simply referred to as an Al crystal grain), a crystal grain of the metal matrix and the Al crystal grain are separated into
20 two phases in the metal matrix. The number average domain diameter of the Al crystal grain is in the range of 0.01 to 500 μm . Hydrogen generating efficiency is influenced by the presence of the Al crystal grain in the metal matrix. When the number average domain diameter of the Al crystal grain is less than 0.01 μm , the hydrogen generating efficiency is low. Thus, this is not industrially practical. When the number average domain diameter of the Al crystal grain exceeds 500 μm , oxidation becomes significant and the capability to generate hydrogen is reduced.

25 **[0013]** Further, the hydrogen gas generating member of the present invention is formed of Al containing unavoidable impurities and a metal X. The metal X is at least one selected from Sn, Bi, and In. Further, the hydrogen gas generating member of the present invention has a composition having a temperature range where the hydrogen gas generating member is separated into Al having a mutually different composition in a liquid phase state and another liquid phase and the two-liquid phase separation is present. Thus, the hydrogen gas generating member of the present invention has
30 a metallic texture which is separated into two-liquid phases of the metal X and Al as the metal matrix.

Further, as for the content of X, X is in the range of Sn: 10.1 to 99.5%, Bi: 30.1 to 99.5%, and In: 10.1 to 99.5% based on the hydrogen gas generating member. When Sn is less than 10.1%, Bi is less than 30.1%, and In is less than 10.1% as to the content, the amount of Al is increased. This provides demerits such that the dispersed region of Al is significantly decreased. When the content exceeds 99.5%, the amount of Al is reduced. This provides demerits such that the generation
35 of hydrogen gas is decreased. Particularly, X is preferably in the range of Sn: 20 to 95%, Bi: 35 to 90%, and In: 25 to 95%. This provides merits such that hydrogen gas is efficiently generated.

Further, the total of two or three metals selected from Sn, Bi, and In as the metal X may be in the range of Sn+Bi: 20.1 to 99.5%, Sn+In: 10 to 99.5%, Bi+In: 20.1 to 99.5%, and Sn+Bi+In: 20 to 99.5%. The two-phase separation into the metal X having a mutually different composition in a liquid phase state and Al is facilitated by containing two or more
40 metals, which makes the production of the alloy easy. When Sn+Bi and Bi+In are less than 20.1%, Sn+In is less than 10%, and Sn+Bi+In is less than 20% as for the content, the amount of Al is increased. This provides demerits such that the dispersed region of Al is significantly decreased. On the other hand, when the content exceeds 99.5%, the amount of Al is reduced. This provides demerits such that the generation of hydrogen gas is decreased. Particularly, the content is preferably in the range of Sn+Bi: 25 to 95%, Sn+In: 25 to 95%, Bi+In: 25 to 95%, and Sn+Bi+In: 25 to 95%. This provides merits such that hydrogen gas is efficiently generated.

45 **[0014]** Further, in the hydrogen gas generating member of the present invention, the Al-X alloy is made to contain Y. Y is at least one selected from Pb: 0.01 to 20%, Zn: 0.01 to 30%, Si: 0.01 to 20%, Cd: 0.01 to 20%, Cu: 0.01 to 20%, Fe: 0.01 to 5%, Ga: 0.01 to 30%, Ge: 0.01 to 30%, Hg: 0.01 to 20%, Mg: 0.01 to 20%, Ni: 0.01 to 5%, S: 0.01 to 5%, and Mn: 0.01 to 5%.

50 Y is an element which has a high tendency to form a liquid two-phase separation system by containing Al. Particularly, when a liquid two-phase separation system Al alloy which contains Sn, Bi, and In is used as a more stable liquid two-phase separation system, it is easier to melt the alloy.

When Y is contained, two-liquid phase separation becomes stable and the melting temperature for the alloy production is reduced. This makes a fine dispersion of the Al crystal grain in the metal matrix easy.

55 Further, the content of Y is within the range of 60% even when a plurality of elements are contained. When the content of Y is less than 0.01%, there are demerits such that the two-liquid phase separation is unstable. When the content of Y exceeds 60%, the amount of Al is reduced. This provides demerits such that the generation of hydrogen gas is decreased. Particularly, it is preferably in the range of 0.1 to 30%. This provides merits such that the production of the

alloy becomes easier. Further, Mn, Zn, Mg, and Ga are particularly preferable as Y. Mn, Zn, and Mg are additive elements of mass-produced aluminum cans. Even if the aluminum cans are melted, hydrogen may be generated. Since the two-liquid phase separation temperature is rapidly decreased by adding Ga, the dissolution of the alloy becomes easier and the hydrogen gas yield becomes larger.

Further, the relation of X and Y is defined by $2X > Y$. When the alloy is produced so that the total of X and Y is in the range of 10.1 to 99.5%, an effect such that a fine crystal grain of Al is dispersed in an X-Y matrix is exerted. Further, when the hydrogen gas generating member contains X, Y, unavoidable impurities, and the balance being Al, the total mass percentage is 100%.

[0015] Further, the hydrogen gas generating member of the present invention is rolled and processed. A rolling apparatus may be any one of a two-high rolling mill, a Sendzimir rolling mill, a planetary rolling mill, and a universal rolling mill. The rolling ratio is preferably in the range of 30 to 99.99%.

The Al crystal grain which is originally fine and has a very high purity in the metal matrix is stretched into a fiber shape through the rolling process. Further, when a large number of fine cracks are formed in the crystal grain, an area coming in contact with water is increased and the reactivity with water may be improved. Further, when the fiber-shaped Al grain is reacted with water, voids are formed in a plate material. When water further permeates into the voids, the whole Al in the metal matrix is reacted with water and thus the hydrogen generating efficiency may be increased.

Further, the hydrogen gas generating member of the present invention is powdered. The powdering may be performed by mechanical grinding processes such as the ball mill grinding process or the attritor grinding process, or, alternatively, the process for powdering from a melted state using, for example, an atomizer. When the hydrogen gas generating member is prepared to have a small particle diameter and further the Al crystal grain is made to be present on the surface, the contact surface with water may be increased and the hydrogen generating efficiency may be increased. At this time, the weight average particle diameter of the hydrogen gas generating member is in the range of 5 to 400 μm . When the weight average particle diameter is less than 5 μm , the surface oxide film of Al is increased. This provides demerits such that the generation of hydrogen is inhibited. When the weight average particle diameter exceeds 400 μm , the dispersed particle diameter of Al becomes larger, which is disadvantageous. Particularly, it is preferably in the range of 20 to 100 μm . This provides merits such that a balance between the dispersed particle diameter of Al and hydrogen generating efficiency is good.

[0016] Further, the process for producing hydrogen gas of the present invention will be described. Fig. 1 is a schematic diagram showing a structure of a hydrogen gas generating apparatus which generates hydrogen gas.

As shown in Fig. 1, a water supply pipe 13 that supplies water and a hydrogen gas collecting pipe 12 that takes the generated hydrogen gas outside to store it are formed in a hydrogen generating apparatus 10. Further, a fixing member 14 for mounting a hydrogen gas generating member 20 is provided in the hydrogen generating apparatus 10 and is brought into contact with a water 15 that is stored inside. Water is supplied from the water supply pipe 13 to the hydrogen generating apparatus 10 and is brought into contact with the hydrogen gas generating member 20. The hydrogen gas generated from the surface is supplied outside through the hydrogen gas collecting pipe 12 and stored in a storage tank (not shown). As shown in Fig. 1, the fixing member 14 not only fixes the hydrogen gas generating member 20 but also may control the hydrogen gas yield when the fixing member 14 is made to be movable up and down and the contact area of the hydrogen gas generating member 20 with water is adjusted. Further, a plurality of the fixed members 14 may be provided and the details will not be described here. The use of the hydrogen generating apparatus 10 having the hydrogen gas generating member 20 allow for easy production of hydrogen gas.

[0017] Subsequently, the hydrogen gas generating member of the present invention will be more specifically described.

(Example 1)

[0018] A process of the rolling treatment of an Al-X-Y alloy of the hydrogen gas generating member will be described. An alloying treatment is performed by first putting Al and Sn as a dopant metal X into the same crucible, melting in a high-frequency melting furnace, and casting to a template. As for the melting method, any of known techniques such as a heating technique using an electric furnace may be used.

The casted Al alloy is subjected to the rolling treatment by rolling. Known materials such as kerosene may be used as lubricants. Further, either hot-rolling or cold-rolling techniques may be used. From a cost viewpoint, the rolling treatment by cold-rolling is preferable.

Fig. 2 is an SEM photograph showing the cross section of the hydrogen gas generating member containing Al and Sn. Al and Sn were cast into a cylindrical template with a diameter of 20 mm after the high-frequency melting. In the SEM photograph, a white base matrix is Sn and a black point is Al. The whole composition analysis was carried out by EPMA. As the result, Al was 30% and Sn was 70%. The specimen was cut out and rolled with a cold rolling mill. The rolling could be easily performed and a hydrogen gas generating member having a rolling ratio of 99.99% could be obtained. The rolled surface of the obtained sample was analyzed XPS (X-ray photoelectron spectroscopy). As the result, it was found that an oxide of Al and an oxide of Sn were present together on the very outermost surface of the rolled sample.

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0.0515 g of hydrogen gas generating member (a thin film) was immersed in water at 15°C. When about 60 seconds passed after the immersion, hydrogen was vigorously generated.

With the expansion of the specimen, a situation where the generation of hydrogen was proceeding from the inside of the specimen was observed. When the generation of hydrogen was finally finished, the shape at the time of rolling of the specimen was not kept. At this time, the hydrogen gas yield was about 6 mL. When the hydrogen gas yield was theoretically, considered, the efficiency was about 90%.

The composition at this time as well as the situation of the generation of hydrogen gas are shown in Table 1. The generation of hydrogen gas was visually evaluated. The case where hydrogen was vigorously generated was defined as "Very good", the case where hydrogen was sufficiently generated in a state in which there was no problem practically was defined as "Good", and the case where hydrogen was not generated or the hydrogen generation level was not enough even if it was generated was defined as "Bad".

(Examples 2 to 19)

[0019] Al-X-Y alloys of Examples 2 to 19 were subjected to rolling treatment in the same manner as described in Example 1 except that the type and amount of the dopant metal as to X and Y were changed. The composition at this time as well as the situation of the generation of hydrogen gas are additionally shown in Table 1.

(Comparative example 1)

[0020] A pure Al was cast into the cylindrical template with a diameter of 20 mm after the high-frequency melting. The specimen was cut out and rolled with the cold rolling mill. A specimen with a rolling ratio of 99.99% was prepared. The specimen was analyzed by XPS. As the result, an oxide of Al was confirmed on the very outermost surface of the sample. Further, the same specimen was brought into contact with water at ordinary temperature. Nothing was generated from the plate material. The results are shown in Table 1.

(Comparative examples 2 to 34)

[0021] Alloys of Comparative examples 2 to 4 to which Sn, Bi, and In were added and alloys of Comparative examples 5 to 11 to which Fe, Zn, Ga, Ge, Hg, Pb, and Si were added were subjected to rolling treatment in the same manner as described in Example 1 except that the type and amount of the dopant metal as to X and Y were changed. The composition at this time as well as the situation of the generation of hydrogen gas are shown in Table 1. Nothing was generated from plate materials of Sn, Bi, and In and other alloys. The results are shown in Table 1.

[0022] The results of XPS of Example 1 and Comparative example 1, it was found that the presence of an oxide of an additive element on the very outermost surface of the sample was important as an initial stage of hydrogen generation. When water penetrates into an interface between the oxide of Al and the oxide of the additive element, hydrogen is generated by reaction with the pure Al therein.

[0023] At this time, the additive element for alloying to Al, the rolling ratio, and the generation of hydrogen were examined. The result is shown in Table 1 including the results of Examples 1 to 34 and Comparative examples 1 to 11.

[Table 1]

Dopant materials to Al-alloy, the rolling ratio, and the generation of hydrogen					
Example Comparative Example	X (mass percent)	Y (mass percent)	Rolling ratio (percent)	Water temperature (°C)	Generation of hydrogen
Example 1	Sn:70	-	99.99	15	Very good
Example 2	Sn:50	Pb:0.01	80	50	Very good
Example 3	Sn:95	-	85	20	Very good
Example 4	Sn:90	-	50	50	Very good
Example 5	Sn:80	-	80	20	Very good
Example 6	Sn:70	-	95	30	Very good
Example 7	Sn:60	-	97	15	Very good
Example 8	Sn:40	-	98	50	Very good

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(continued)

Dopant materials to Al-alloy, the rolling ratio, and the generation of hydrogen					
Example Comparative Example	X (mass percent)	Y (mass percent)	Rolling ratio (percent)	Water temperature (°C)	Generation of hydrogen
Example 9	Sn:30	-	98	90	Very good
Example 10	Sn:20	-	99	70	Very good
Example 11	Sn:5	-	99	50	Good
Example 12	Sn:20	Si:20	15	90	Good
Example 13	Bi:90	-	10	50	Very good
Example 15	Bi:50	-	10	20	Very good
Example 16	Bi:20	-	50	90	Good
Example 17	Bi:50	Cd:10	30	30	Very good
Example 18	Bi:30.1	Cu:0.2	30	90	Good
Example 19	In:90	-	90	15	Very good
Example 20	In:30	-	90	15	Very good
Example 21	Sn:48, Bi:20	-	95	35	Very good
Example 22	Sn:48, Bi:20	Fe:3	85	15	Very good
Example 23	Sn:57, Bi:5	Zn:0.3	95	70	Very good
Example 24	Sn:48, Bi:20	Ga:1	80	90	Very good
Example 25	Sn:16, Bi:20	Ge:25	40	30	Good
Example 26	Sn:19, Bi:5	Hg:0.2	90	90	Good
Example 27	Sn:47.5, Bi:5	Mg:10	99	60	Good
Example 28	Sn:40, In:40	-	90	15	Very good
Example 29	Sn:40, In:40	Ni:0.1	90	15	Very good
Example 30	Sn:50, In:5	S:0.01	99	60	Very good
Example 31	Sn:3, Bi:7, In:10	Mn:1	50	50	Very good
Example 32	Sn:6, Bi:14, In:20	-	15	90	Very good
Example 33	Sn:4, In:5, Bi:11	-	80	15	Very good
Example 34	Sn:7, In:10, Bi:22	-	30	90	Very good
Comparative Example 1	-	-	99.99	90	Bad
Comparative Example 2	Sn:0.01	-	99.99	80	Bad
Comparative Example 3	Bi:0.01	-	99.99	90	Bad
Comparative Example 4	In:0.01	-	99.99	70	Bad
Comparative Example 5	-	Fe:4	70	80	Bad

(continued)

Dopant materials to Al-alloy, the rolling ratio, and the generation of hydrogen					
Example Comparative Example	X (mass percent)	Y (mass percent)	Rolling ratio (percent)	Water temperature (°C)	Generation of hydrogen
Comparative Example 6	-	Zn:3	99	80	Bad
Comparative Example 7	-	Ga:15	70	80	Bad
Comparative Example 8	-	Ge:10	95	80	Bad
Comparative Example 9	-	Hg:5	50	80	Bad
Comparative Example 10	-	Pb:10	95	90	Bad
Comparative Example 11	-	Si:10	40	80	Bad

[0024] As is apparent from Table 1, it was found that hydrogen gas could be generated by adding Sn, Bi, and In as additive elements to Al. Further, the hydrogen generating efficiency was higher as the rolling ratio was higher.

(Example 35)

[0025] A process of powdering the Al-X-Y alloy of the hydrogen gas generating member of the present invention will be described.

First, the Al-X-Y alloy is put into the same crucible and is melted in the high-frequency melting furnace, and then it is formed into powder using powder forming techniques such as the gas atomizing method. As for the melting method, any of known techniques such as a heating technique using an electric furnace may be used.

The Al-X-Y alloy was formed into powder by the gas atomizing method after the high-frequency melting. The composition of the produced powder is shown in Table 2.

Fig. 3 is an SEM photograph showing the cross section of the hydrogen gas generating member containing Al, Sn, and In. A white base matrix is Sn-In and a black point is Al. The whole composition analysis was carried out by EPMA. As the result, Al was 60% and Sn-In alloy was 40%.

5 g of the powder was taken out and stored in a stainless steel mesh, followed by bringing into contact with warm water at 60°C. The water discolored while hydrogen was vigorously generated immediately after the immersion. An SEM photograph the cross-sectional texture of the test powder at the time is shown in Fig. 4. A crack entered from the surface and extended to the inside of Al. In this case, the active surface of Al was exposed. It is considered that hydrogen was generated by the hydration reaction of hydroxide formation with water. The composition at this time as well as the situation of the generation of hydrogen gas are shown in Table 2.

(Examples 36 to 53)

[0026] Al-X-Y alloys of Examples 36 to 53 were subjected to powdering treatment in the same manner as described in Example 35 except that the type and amount of the dopant metal as to X and Y were changed. The composition at this time as well as the situation of the generation of hydrogen gas are additionally shown in Table 2.

(Comparative examples 12 to 18)

[0027] Al-X-Y alloys of Comparative examples 12 to 18 were subjected to powdering treatment in the same manner as described in Example 35 except that the type and amount of the dopant metal as to X and Y were changed. The composition at this time as well as the situation of the generation of hydrogen gas are additionally shown in Table 2. 5 g of the powder was taken out, followed by bringing into contact with warm water. No change occurred after the immersion and the generation of hydrogen was not observed.

[0028] At this time, the additive element for alloying to Al, the particle diameter, and the generation of hydrogen were

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examined. The result is shown in Table 2 including the results of Examples 35 to 53 and Comparative examples 12 to 18.

[Table 2]

Dopant materials to Al-alloy, the particle diameter, and the generation of hydrogen					
Example Comparative Example	X(mass percent)	Y(mass percent)	Particle diameter of powder (μm)	Water temperature (°C)	Generation of hydrogen
Example 35	Sn:30, In:10	-	200-250	60	Very good
Example 36	Sn:90	-	10-50	90	Very good
Example 37	Sn:40	Pb:0.01	50-200	70	Very good
Example 38	Sn:20	Si:20	10-250	90	Good
Example 39	Bi:99.9	-	5-100	50	Good
Example 40	Bi:50	Cd:10	100-300	30	Very good
Example 41	Bi:30.1	Cu:0.2	200-250	90	Good
Example 42	In:90	-	10-100	15	Very good
Example 43	Sn:30, Bi:5	Fe:3	10-150	50	Very good
Example 44	Sn:40, Bi:5	Zn:3	50-150	70	Very good
Example 45	Sn:48, Bi:20	Ga:1	10-250	90	Very good
Example 46	Sn:36, Bi:20	Ge:25	50-200	30	Good
Example 47	Sn:19, Bi:5	Hg:0.2	10-250	90	Good
Example 48	Sn:47.5, Bi:5	Mg:10	5-100	60	Good
Example 49	Sn:15, In:5	Ni:0.1	100-300	90	Very good
Example 50	Sn:13, Bi:7, In: 30	S:0.01	10-100	50	Very good
Example 51	Sn:22, Bi:4, In: 35	Mn:0.2	10-150	90	Very good
Example 52	Sn:37, In:5, Bi: 30	-	200-250	15	Very good
Example 53	Sn:7, In:10, Bi: 22	-	10-100	90	Very good
Comparative Example 12	-	Cu:50	30-400	90	Bad
Comparative Example 13	Sn:0.01	-	10-100	80	Bad
Comparative Example 14	Bn:0.01	-	50-300	90	Bad
Comparative Example 15	In:0.01	-	30-400	70	Bad
Comparative Example 16	-	Fe:4	200-250	80	Bad
Comparative Example 17	-	Pb:20	5-100	80	Bad
Comparative Example 18	-	Ga:15	50-100	80	Bad

[0029] As is apparent from Table 2, Sn, In, and Bi were good as additive elements for alloying to Al. Further, the hydrogen generating efficiency was higher as the particle diameter was smaller.

Industrial Applicability

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[0030] A lot of applications such as portable batteries and emergency electrical generators are considered using the hydrogen gas generating member and the process for producing hydrogen gas to be obtained by the present invention. Further, it is a safe material as science teaching material to understand the mechanism of water.

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Claims

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1. A hydrogen gas generating member comprising: a texture in which Al is finely dispersed in a metal matrix, wherein hydrogen gas is generated by bringing the hydrogen gas generating member into contact with water.

2. The hydrogen gas generating member according to claim 1, wherein the hydrogen gas generating member is separated into Al having a mutually different composition in a liquid phase state and another liquid phase and the two-liquid phase separation is present.

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3. The hydrogen gas generating member according to claim 1 or 2, wherein the hydrogen gas generating member is an Al-X alloy and X is at least one selected from Sn, Bi, and In.

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4. The hydrogen gas generating member according to claim 3, wherein the hydrogen gas generating member is an Al-X alloy and X is selected from Sn: 10.1 to 99.5% (mass percent: hereinafter on the basis of mass percent), Bi: 30.1 to 99.5%, In: 10.1 to 99.5%, Sn+Bi: 20.1 to 99.5%, Sn+In: 10 to 99.5%, Bi+In: 20.1 to 99.5%, and Sn+Bi+In: 20 to 99.5%.

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5. The hydrogen gas generating member according to claim 3 or 4, wherein the Al-X alloy further includes Y and Y is at least one selected from Pb: 0.01 to 20%, Zn: 0.01 to 30%, Si: 0.01 to 20%, Cd: 0.01 to 20%, Cu: 0.01 to 20%, Fe: 0.01 to 5%, Ga: 0.01 to 30%, Ge: 0.01 to 30%, Hg: 0.01 to 20%, Mg: 0.01 to 20%, Ni: 0.01 to 5%, S: 0.01 to 5%, and Mn: 0.01 to 5%, the content is within the range of 60% even when a plurality of elements are contained, and the relation of X and Y is defined by $2X > Y$.

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6. The hydrogen gas generating member according to any one of claims 1 to 5, wherein the hydrogen gas generating member is rolled and processed.

7. The hydrogen gas generating member according to any one of claims 1 to 5, wherein the hydrogen gas generating member is powdered.

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8. A process for producing hydrogen gas comprising a step of: bringing the hydrogen gas generating member according to any one of claims 1 to 7 into contact with water to generate hydrogen gas.

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Fig. 1

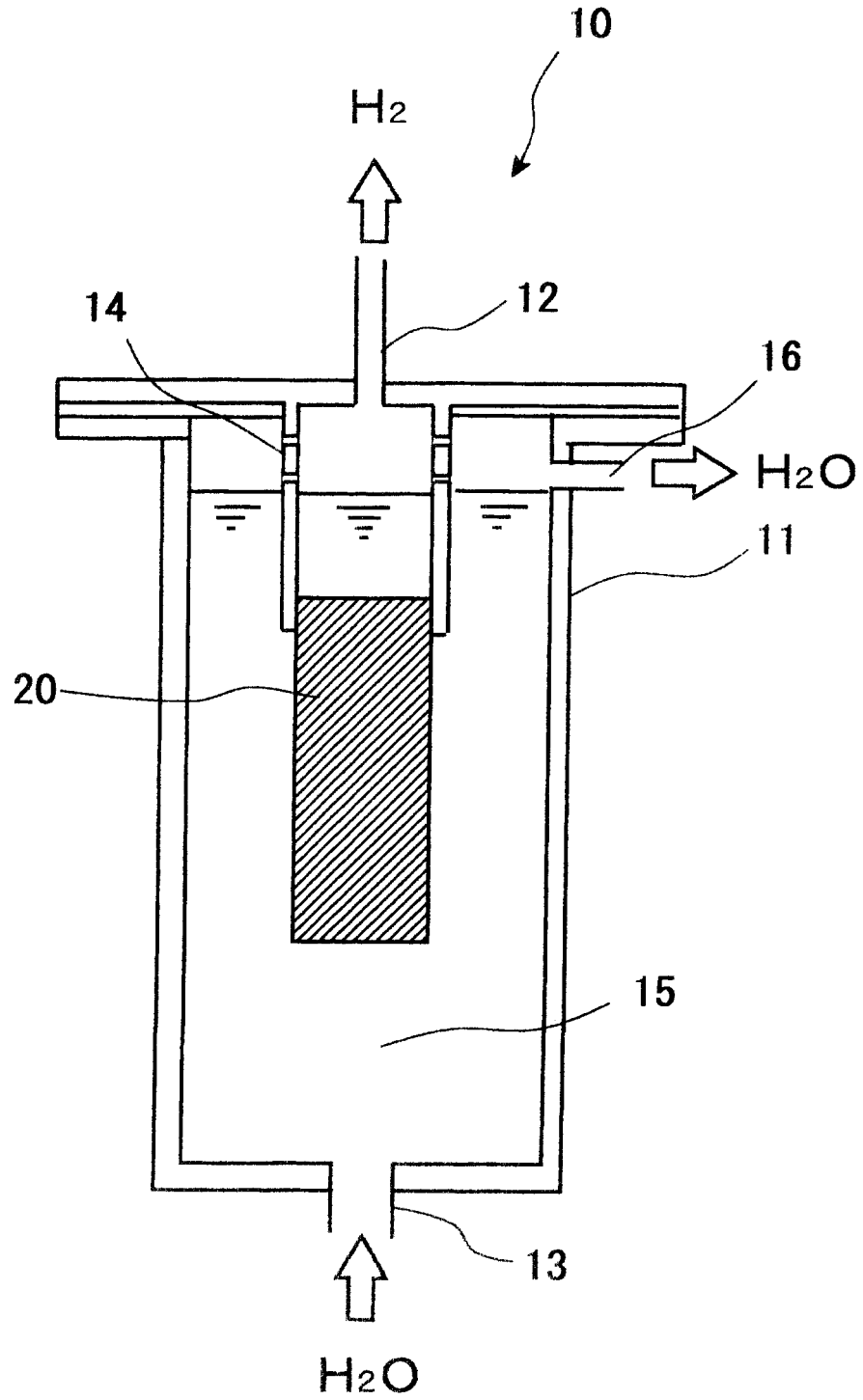


Fig. 2

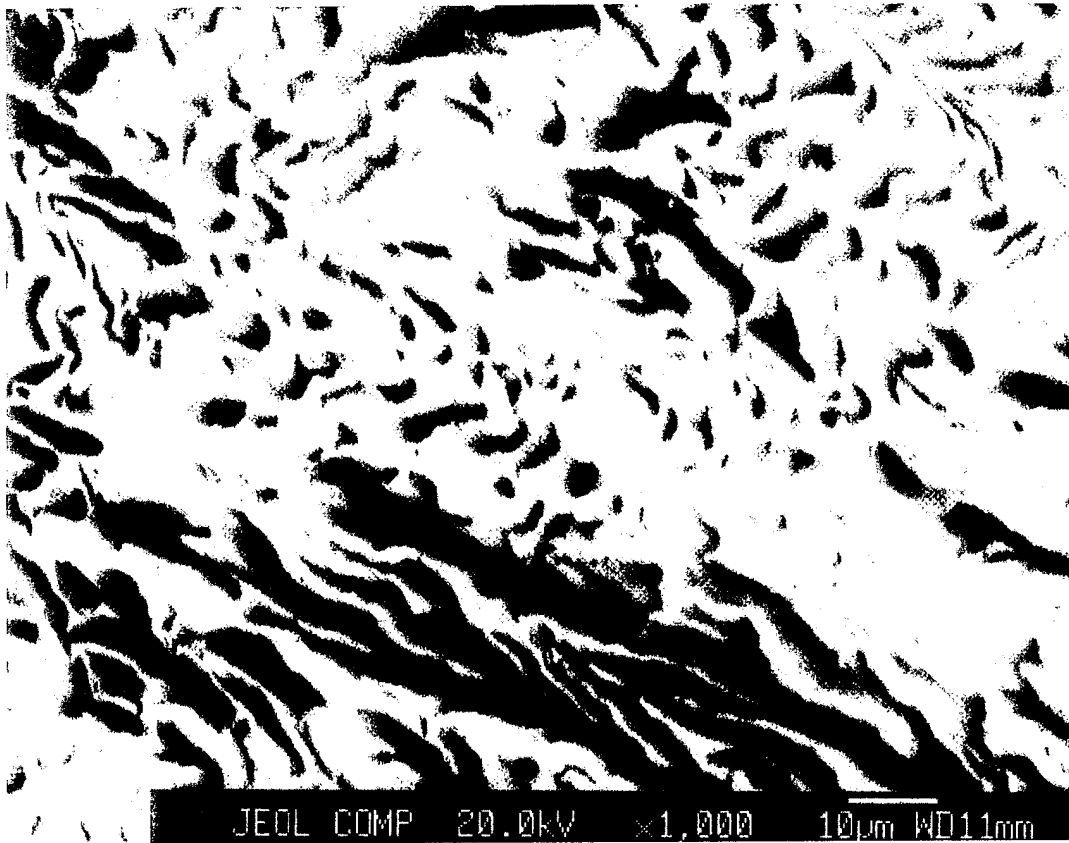


Fig. 3

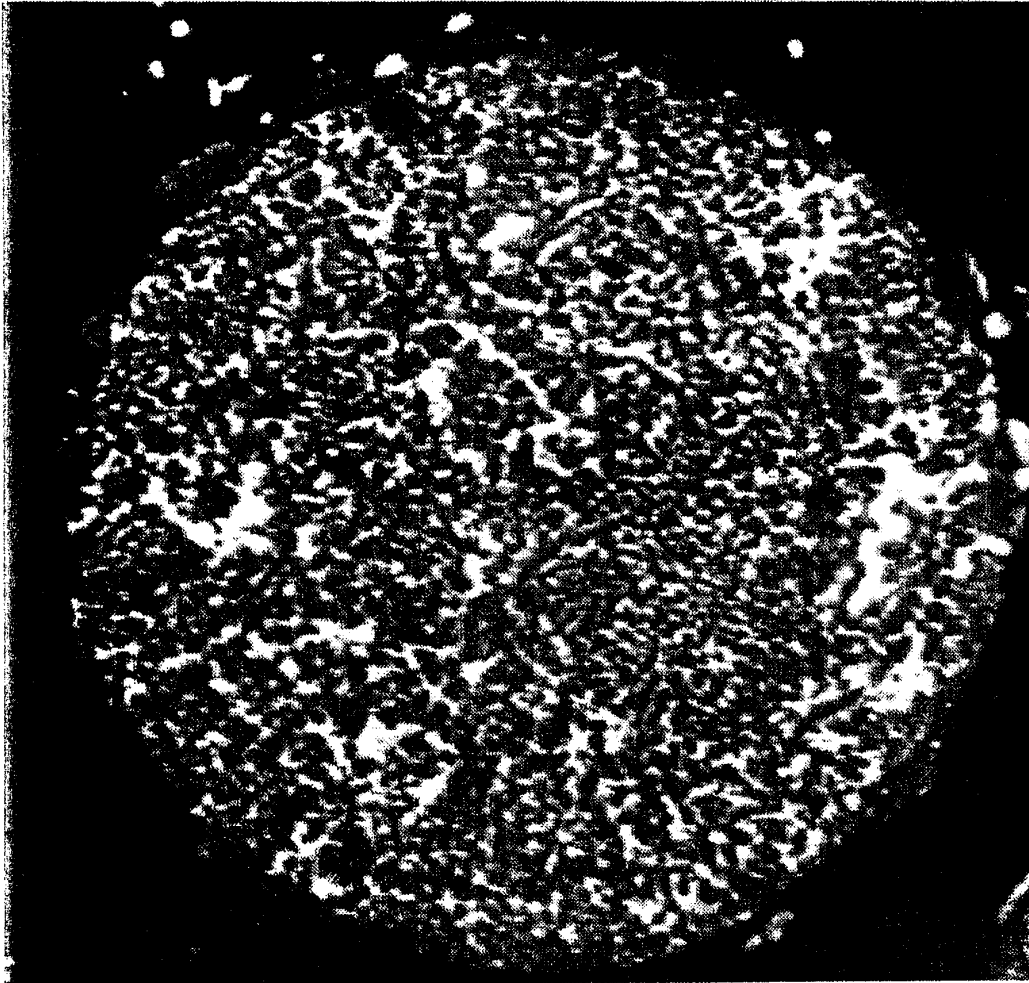
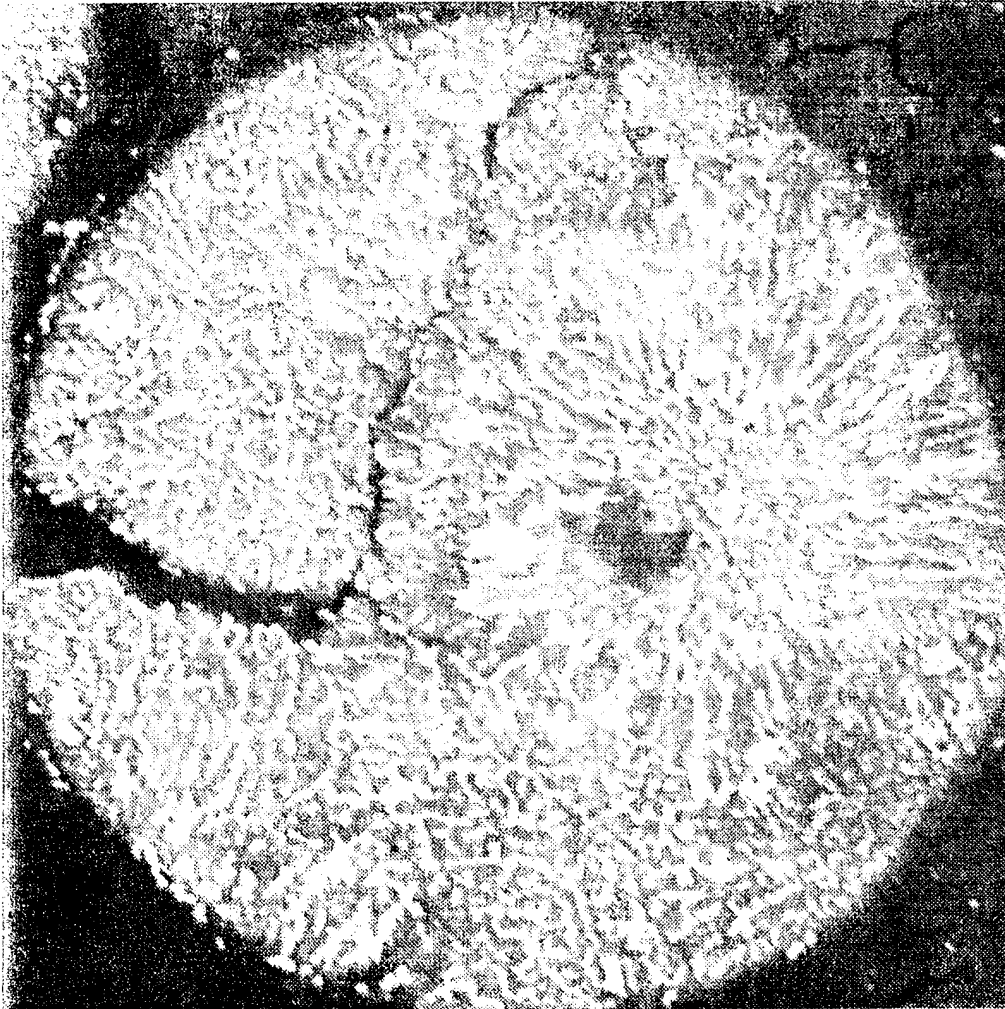


Fig. 4



INTERNATIONAL SEARCH REPORT

International application No.
PCT/JP2008/002191

<p>A. CLASSIFICATION OF SUBJECT MATTER C01B3/08(2006.01)i, B22F1/00(2006.01)i, C22C12/00(2006.01)i, C22C13/00(2006.01)i, C22C21/00(2006.01)i, C22C30/04(2006.01)i</p> <p>According to International Patent Classification (IPC) or to both national classification and IPC</p>														
<p>B. FIELDS SEARCHED</p> <p>Minimum documentation searched (classification system followed by classification symbols) C01B3/08, B22F1/00, C22C12/00, C22C13/00, C22C21/00, C22C30/04</p> <p>Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched Jitsuyo Shinan Koho 1922-1996 Jitsuyo Shinan Toroku Koho 1996-2008 Kokai Jitsuyo Shinan Koho 1971-2008 Toroku Jitsuyo Shinan Koho 1994-2008</p> <p>Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) Science Direct</p>														
<p>C. DOCUMENTS CONSIDERED TO BE RELEVANT</p> <table border="1"> <thead> <tr> <th>Category*</th> <th>Citation of document, with indication, where appropriate, of the relevant passages</th> <th>Relevant to claim No.</th> </tr> </thead> <tbody> <tr> <td>X</td> <td>JP 1-23543 B2 (Osaka Fuji Corp.), 02 May, 1989 (02.05.89), Claims; page 1, left column, line 24 to page 3, left column, line 8 & US 4752463 A & EP 248960 A1</td> <td>1-8</td> </tr> <tr> <td>X A</td> <td>Mei-Qiang Fan et al., Studies on hydrogen generation characteristics of hydrolysis of the ball milling Al-based materials in pure water, International Journal of Hydrogen Energy, 2007.02.17, vol.32, No.14, pp.2809- 2815</td> <td>1-4, 7, 8 5, 6</td> </tr> <tr> <td>A</td> <td>JP 2002-161325 A (Ulvac, Inc.), 04 June, 2002 (04.06.02), Claims (Family: none)</td> <td>1-8</td> </tr> </tbody> </table>			Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.	X	JP 1-23543 B2 (Osaka Fuji Corp.), 02 May, 1989 (02.05.89), Claims; page 1, left column, line 24 to page 3, left column, line 8 & US 4752463 A & EP 248960 A1	1-8	X A	Mei-Qiang Fan et al., Studies on hydrogen generation characteristics of hydrolysis of the ball milling Al-based materials in pure water, International Journal of Hydrogen Energy, 2007.02.17, vol.32, No.14, pp.2809- 2815	1-4, 7, 8 5, 6	A	JP 2002-161325 A (Ulvac, Inc.), 04 June, 2002 (04.06.02), Claims (Family: none)	1-8
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<p>Date of the actual completion of the international search 25 August, 2008 (25.08.08)</p>		<p>Date of mailing of the international search report 02 September, 2008 (02.09.08)</p>												
<p>Name and mailing address of the ISA/ Japanese Patent Office</p>		<p>Authorized officer</p>												
<p>Facsimile No.</p>		<p>Telephone No.</p>												

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- **O.V Kravchenko et al.** *Journal of Alloys and Compounds*, 2005, vol. 397, 58-62 [0004]
- **M-Q.Fan et al.** *International Journal of Hydrogen Energy*, *Corrected Proof*, 12 February 2007 [0004]